

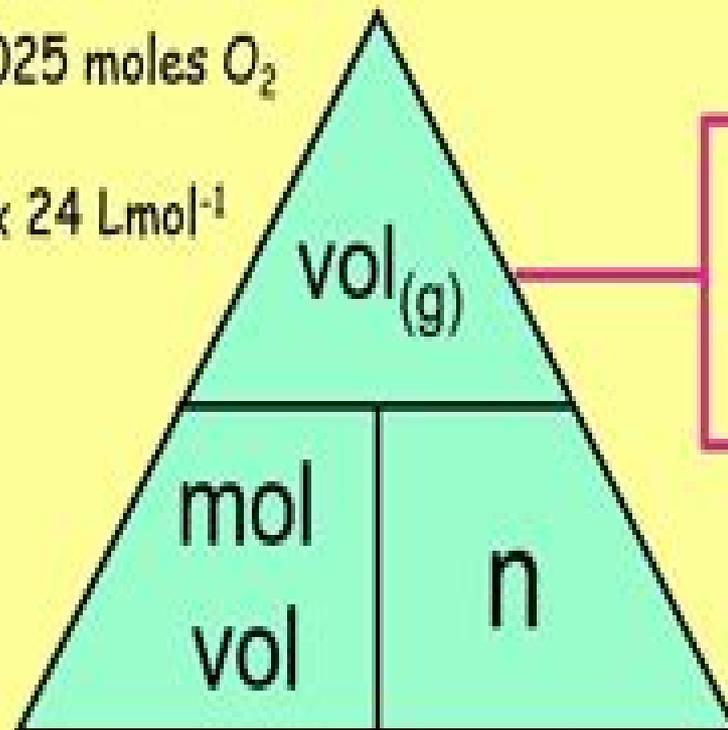
Assume molar volume is 24Lmol^{-1}

Calculate volume of 0.025 moles O_2

$$\text{Vol} = 0.025 \text{ mol} \times 24 \text{ Lmol}^{-1}$$

$$= 0.6 \text{ L}$$

$$= 600 \text{ cm}^3$$



Gas Volume

$$\text{Vol} = n \times \text{mol.vol}$$

Calculations Using Standard Molar Volume

Zhe Xu



Calculations Using Standard Molar Volume:

Vacuum and Ultravacuum Igor Bello, 2017-11-02 Vacuum technology has enormous impact on human life in many aspects and fields such as metallurgy material development and production food and electronic industry microelectronics device fabrication physics materials science space science engineering chemistry technology of low temperature pharmaceutical industry and biology All decorative coatings used in jewelries and various daily products including shiny decorative papers the surface finish of watches and light fixtures are made using vacuum technological processes Vacuum analytical techniques and vacuum technologies are pillars of the technological processes material synthesis deposition and material analyses all of which are used in the development of novel materials increasing the value of industrial products controlling the technological processes and ensuring the high product quality Based on physical models and calculated examples the book provides a deeper look inside the vacuum physics and technology *Physical Chemistry* Robert J. Silbey, Robert A. Alberty, George A. Papadantonakis, Mounji G. Bawendi, 2022-06-15 Ever since *Physical Chemistry* was first published in 1913 it has remained a highly effective and relevant learning tool thanks to the efforts of physical chemists from all over the world Each new edition has benefited from their suggestions and expert advice The result of this remarkable tradition is now in your hands

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new applications and insights as well as new Mathematical Background inter chapters to help students review key quantitative concepts This is a splendid book True to the authors philosophy as outlined in the preface it approaches physical chemistry by first developing the quantum theory of molecular electronic structure then by statistical arguments moves into thermodynamics and thence to kinetics Peter Taylor Review in Chemistry World Royal Society of Chemistry July 31 2009

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