

## Acid-Base Titration Lab Report Sheet

Name:	Quiz is 20 points
	Lab Report is 80 points

### Part I: Standardization of NaOH using Anhydrous Citric Acid

	Trial 1	Trial 2	Trial 3
Starting M of citric acid solution	$8.328 \times 10^{-3} \text{ M}$	$8.328 \times 10^{-3} \text{ M}$	$8.328 \times 10^{-3} \text{ M}$
Volume of citric acid	20.00 mL	20.00 mL	20.00 mL
Moles of citric acid	$1.667 \times 10^{-4} \text{ mol}$	$1.667 \times 10^{-4} \text{ mol}$	$1.667 \times 10^{-4} \text{ mol}$
Moles of NaOH	$4.997 \times 10^{-4} \text{ mol}$	$4.997 \times 10^{-4} \text{ mol}$	$4.997 \times 10^{-4} \text{ mol}$
Initial graduated cylinder volume	10.0 mL	10.0 mL	10.0 mL
Final graduated cylinder volume	5.2 mL	5.4 mL	5.3 mL
Volume of NaOH used	4.8 mL	4.6 mL	4.7 mL
Volume of NaOH used in L	0.0048 L	0.0046 L	0.0047 L
NaOH concentration	0.104 M	0.109 M	0.106 M

Average Molarity of NaOH \_\_\_\_\_ 0.106 M \_\_\_\_\_

Show your work for one trial as well as the average molarity.

# Acid Base Titration Lab Report Answers Chemfax

**Thomas Griffiths**



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## **Table of Contents Acid Base Titration Lab Report Answers Chemfax**

1. Understanding the eBook Acid Base Titration Lab Report Answers Chemfax
  - The Rise of Digital Reading Acid Base Titration Lab Report Answers Chemfax
  - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titration Lab Report Answers Chemfax
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acid Base Titration Lab Report Answers Chemfax
  - User-Friendly Interface

4. Exploring eBook Recommendations from Acid Base Titration Lab Report Answers Chemfax
  - Personalized Recommendations
  - Acid Base Titration Lab Report Answers Chemfax User Reviews and Ratings
  - Acid Base Titration Lab Report Answers Chemfax and Bestseller Lists
5. Accessing Acid Base Titration Lab Report Answers Chemfax Free and Paid eBooks
  - Acid Base Titration Lab Report Answers Chemfax Public Domain eBooks
  - Acid Base Titration Lab Report Answers Chemfax eBook Subscription Services
  - Acid Base Titration Lab Report Answers Chemfax Budget-Friendly Options
6. Navigating Acid Base Titration Lab Report Answers Chemfax eBook Formats
  - ePub, PDF, MOBI, and More
  - Acid Base Titration Lab Report Answers Chemfax Compatibility with Devices
  - Acid Base Titration Lab Report Answers Chemfax Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acid Base Titration Lab Report Answers Chemfax
  - Highlighting and Note-Taking Acid Base Titration Lab Report Answers Chemfax
  - Interactive Elements Acid Base Titration Lab Report Answers Chemfax
8. Staying Engaged with Acid Base Titration Lab Report Answers Chemfax
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acid Base Titration Lab Report Answers Chemfax
9. Balancing eBooks and Physical Books Acid Base Titration Lab Report Answers Chemfax
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acid Base Titration Lab Report Answers Chemfax
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Acid Base Titration Lab Report Answers Chemfax
  - Setting Reading Goals Acid Base Titration Lab Report Answers Chemfax
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Acid Base Titration Lab Report Answers Chemfax
  - Fact-Checking eBook Content of Acid Base Titration Lab Report Answers Chemfax
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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