

Atomic Orbitals

Use with Chapter 5,
Section 5.2

1. What is the shape of an s orbital?
Spherical
2. What is the relationship between the size of an s orbital and the principal energy level in which it is found?

3. What is the shape of a p orbital? How many p orbitals are there in a sublevel?

4. How many electrons can each orbital hold?
2
5. Look at the diagrams of the p orbitals. What do x, y, and z refer to?

6. How many d orbitals are there in a given sublevel? How many total electrons can the d orbitals in a sublevel hold?
5 10
7. Which d orbitals have the same shape?
 $x^2 - y^2$
8. What point in each diagram represents an atom's nucleus?

9. How likely is it that an electron occupying a p or a d orbital would be found very near an atom's nucleus? What part of the diagram supports your conclusion?

Teaching Transparency 16 Answers

K Payea



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Reviewing **Teaching Transparency 16 Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Teaching Transparency 16 Answers**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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