

# UNIT 5 COORDINATION COMPOUNDS

## Multiple Choice Questions (MCQs):

- A purple compound 'M' has a formula  $\text{Co}(\text{NH}_3)_6\text{Cl}_2$ , on adding excess  $\text{AgNO}_3$  to one mole of 'M' produces 2 moles of  $\text{AgCl}$ . If secondary valence of Co is 6 then correct formula of the complex will be?
  - $[\text{CoCl}_2(\text{NH}_3)_6]$
  - $[\text{Co}(\text{NH}_3)_6]\text{Cl}_2$
  - $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2 \cdot \text{NH}_3$
  - $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2$
- Which among the following is not a double salt?
  - Mohr's Salt
  - Potash Alum
  - Carnallite
  - Potassium ferrocyanide.
- The correct order of the stoichiometries of  $\text{AgCl}$  formed when  $\text{AgNO}_3$  in excess is treated with the complexes:  $\text{CoCl}_2 \cdot 6\text{NH}_3$ ,  $\text{CoCl}_2 \cdot 5\text{NH}_3$ ,  $\text{CoCl}_2 \cdot 4\text{NH}_3$  respectively is (coordination valency is 6):
  - 3  $\text{AgCl}$ , 1  $\text{AgCl}$ , 2  $\text{AgCl}$
  - 3  $\text{AgCl}$ , 2  $\text{AgCl}$ , 1  $\text{AgCl}$
  - 2  $\text{AgCl}$ , 3  $\text{AgCl}$ , 1  $\text{AgCl}$
  - 1  $\text{AgCl}$ , 3  $\text{AgCl}$ , 2  $\text{AgCl}$
- Which among the following has highest molar conductivity?
  - $\text{K}_4[\text{Fe}(\text{CN})_6]$
  - $[\text{Co}(\text{NH}_3)_6]\text{Cl}_2$
  - $[\text{Cu}(\text{NH}_3)_4]\text{SO}_4$
  - $[\text{Ni}(\text{CO})_4]$
- The primary and secondary valencies of  $[\text{Fe}(\text{C}_2\text{O}_4)_3]^{3-}$  complexes are
  - 3 and 3
  - 3 and 6
  - 6 and 3
  - 0 and 3
- Coordination number of the central atom/ion is determined by
  - only by the number of pi bonds formed by the ligand with the central atom/ion.
  - both by the number of sigma and pi bonds formed by the ligand with the central atom/ion.
  - only by the number of sigma bonds formed by the ligand with the central atom/ion.
  - by the total number of electrons present in ligand.
- The correct IUPAC name of  $\text{K}_3[\text{Cr}(\text{C}_2\text{O}_4)_3]$  is
  - Potassiumtrioxalatochromium(III)
  - Potassiumtrioxalatochromate(III)
  - Potassiumtrisoxalatochromate(III)
  - Potassium oxalatochromate(III)
- $[\text{Pt}(\text{NH}_3)_4\text{Cl}_2] \text{Br}_2$  complex can show :
  - Hydrated as well as ionization isomerism
  - Ionization as well as geometrical isomerism
  - Linkage as well as geometrical isomerism
  - Ionization as well as optical isomerism
- Possible isomerism in complexes  $[\text{Co}(\text{NH}_3)_5(\text{NO}_2)]$  and  $[\text{Co}(\text{NH}_3)_5(\text{NO}_2)]\text{Cl}_2$  respectively are:
  - Linkage and optical
  - Geometrical and linkage
  - Optical and ionization
  - Linkage and geometrical
- In the reactions  $[\text{CoCl}_2(\text{NH}_3)_4]^{+} + \text{Cl}^{-} \rightarrow [\text{CoCl}_3(\text{NH}_3)_3] + \text{NH}_3$  two isomers of the product are obtained. The initial complex is
  - cis* isomer
  - trans* isomer
  - cis* or *trans* isomers
  - None of these

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