

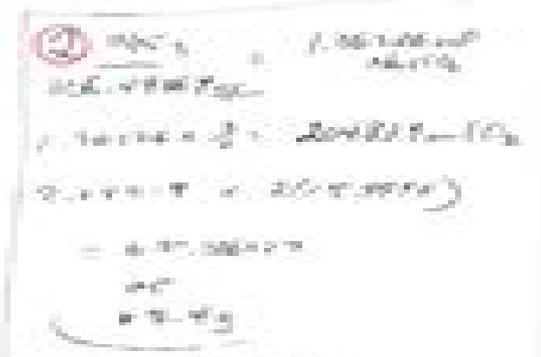
KEY

1. Complete and classify the following partial equations

- a)  $1 \text{ Fe} + 3 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ Fe SO}_4 + 3 \text{ H}_2$  (single repl)
- b)  $\text{NaI} + \text{Cl}_2 \rightarrow$
- c)  $1 \text{ Zn} + 1 \text{ Pb(NO}_3)_2 \rightarrow 1 \text{ Zn(NO}_3)_2 + 1 \text{ Pb}$  (single repl)
- d)  $\text{Al(OH)}_3 + \text{HCl} \rightarrow$
- e)  $2 \text{ Cu} + \text{O}_2 \rightarrow 2 \text{ CuO}$        $\text{4 Cu} + 1 \text{ O}_2 \rightarrow 2 \text{ Cu}_2\text{O}$  (synthesis)
- f)  $\text{Al}_2\text{O}_3 \rightarrow$
- g)  $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$  (synth)
- h)  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- i)  $1 \text{ Mg} + 2 \text{ HNO}_3 \rightarrow 1 \text{ Mg(NO}_3)_2 + 1 \text{ H}_2$  (s.r.)
- j)  $\text{Na} + \text{H}_2\text{O} \rightarrow$
- k)  $3 \text{ NaOH} + \text{H}_3\text{PO}_4 \rightarrow 1 \text{ Na}_3\text{PO}_4 + 3 \text{ H}_2\text{O}$  (acid base)
- l)  $\text{H}_2\text{SO}_4 + \text{Al(OH)}_3 \rightarrow$
- m)  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$  (combustion)
- n)  $\text{AlCl}_3 \rightarrow$
- o)  $2 \text{ KOH} + \text{H}_2\text{SO}_4 \rightarrow 1 \text{ K}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$  (acid base)
- p)  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- q)  $1 \text{ NaOH} + \text{HCl} \rightarrow 1 \text{ NaCl} + 1 \text{ H}_2\text{O}$  (acid base)
- r)  $\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow$
- s)  $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$  (combustion)
- t)  $\text{H}_2\text{O} \rightarrow$
- u)  $2 \text{ HNO}_3 + \text{CuCO}_3 \rightarrow 1 \text{ Cu(NO}_3)_2 + 1 \text{ CO}_2$  (s.r.)
- v)  $\text{CH}_4 + \text{O}_2 \rightarrow$
- w)  $1 \text{ Pb} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ PbSO}_4 + 1 \text{ H}_2$  (s.r.)
- x)  $\text{KI} + \text{Br}_2 \rightarrow$
- y)  $1 \text{ Ba(OH)}_2 + 2 \text{ HNO}_3 \rightarrow 1 \text{ Ba(NO}_3)_2 + 2 \text{ H}_2\text{O}$  (acid base)
- z)  $\text{Pb} + \text{Cu(NO}_3)_2 \rightarrow$



1  
2  
3  
4  
5



- 2. Given:  $2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$  Calculate the mass of oxygen gas produced if 1.25 g of  $\text{NaClO}_3$  is decomposed. **65.4g**
- 3. Calculate the mass of oxygen gas required to burn 242 g of wax ( $\text{C}_{21}\text{H}_{44}$ ). **254g**
- 4. If 1.40 moles of butane ( $\text{C}_4\text{H}_{10}$ ) is burned
  - a) calculate the number of moles of oxygen gas required. **9.8 moles**
  - b) calculate the number of moles of carbon dioxide produced. **5.6 moles CO<sub>2</sub>**
  - c) calculate the mass of water produced. **10.8 moles H<sub>2</sub>O**
- 5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.
  - a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. **27.4g Al**
  - b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. **192 g Zn**
- 6. Given:  $\text{Cu} + 2 \text{ AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2 \text{ Ag}$ 
  - a) How many grams of Cu are required to react with 3.00 g of silver nitrate? **0.544g Cu**
  - b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? **1.70g Ag**

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