

A Sample Double-Entry Journal

Text Source: "Scientists Fear Spread of Exotic Snakes"

On The Left: The Source Direct Quote and Page Number	On The Right: The Response Sample Responses
<p>"Scientists fear that Burmese pythons, already known to be breeding in South Florida, could spread through much of the southern United States" (UPI 151).</p>	<p>Questions How did these snakes end up in the wild in Florida?</p>
<p>"The snakes weigh up to 250 pounds and slither at a rate of 20 miles per month" (UPI 151).</p> <p>"If pet pythons were introduced into the wild in California by irresponsible pet owners, as happened in Florida, they could become established here even faster, without need of a cross country journey" (UPI 151).</p> <p>"So the absence of major predators outside of Florida could help the snakes on their journey west" (UPI 151).</p>	<p>Summary</p> <p>The Burmese pythons have the potential to spread through the southern part of the United States very rapidly. The snakes have the ability to move many miles from their breeding ground, and since they do not have any major predators the number of snakes will quickly increase.</p>
<p>"In Florida, they eat bobcats, deer, alligators, raccoons, cats, rats, rabbits, muskrats, possum, mice, ducks, egrets, herons and song birds" (UPI 151).</p>	<p>Reflections If the Burmese pythons spread they may wipe out other populations of animals in some areas.</p>

Double Entry Journal Wikispaces

Jianjun Gao



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web limiting reactant theoretical and percentage yields key $4 \text{KO}_2 + 2 \text{H}_2\text{O} \rightarrow 4 \text{KOH} + 3 \text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{CuNO}_3 + 2\text{Ag} + 2\text{AgNO}_3$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with $8.3 \text{ mole H}_2\text{S}$ and multiply it with $2 \text{ mole H}_2\text{O}$ over $2 \text{ mole H}_2\text{S}$ from the equation $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
 $8.3 \text{ mole H}_2\text{O} \times 18 \text{ g H}_2\text{O} = 149.4 \text{ g H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g H}_2\text{O}}{149.4 \text{ g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts/html/limiting_reagents_theoretical_actual_and_percent_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web percent yield is a measure of the efficiency of the experimental design yield efficiency $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$
 b determine the theoretical yield of Al_2O_3
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$
 c determine the percent yield
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$
 theoretical 302.3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333 8 g I_2O_5 1 mol I_2O_5 28 g co 1 mol co

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