

# PHYSICAL PHARMACY

**PHYSICAL CHEMISTRY** – branch of chemistry concerned with physical properties and structure of matter and with laws and theories of physical and chemical changes

**PHYSICAL PHARMACY** – application of physical chemistry in pharmacy; study of physicochemical properties of substances used in drug formulation

## I. FORCES OF ATTRACTION

### A. INTRAMOLECULAR – intra (within) the molecule

- **Ionic Bond** – transfer of electrons; metal (electron donor) + non-metal (electron acceptor); (ex. Na:  $p^+ = 11$ ;  $e^- = 11 \rightarrow p^+ = 11$ ;  $e^- = 10 \rightarrow$  charge = +1)

#### ○ PERIODIC TRENDS REPRESENTING BONDS

- **Ionization Energy** – energy required remove electron (low M; high NM)
- **Electronegativity** – ability to attract electrons (low M; high NM)

- **Covalent Bonds** – sharing of electrons; non-metal + non-metal; polarity – difference in electronegativity (symmetrical = non-polar)

- **Polar** – unequal sharing of electrons; dipole moment – (ex.  $H^{\delta+}-Br^{\delta-}$ )
- **Non Polar** – equal sharing of electrons (ex.  $Br-Br$ )

### B. INTERMOLECULAR – inter (between) the molecules

- **Van der Waals**

- **Dipole-Dipole**  $\rightarrow P + P$ ; Keesom Forces; orientation effect
- **Dipole-Induced Dipole**  $\rightarrow P + NP$ ; Debye Forces; inductive effect
- **Induced Dipole-Induced Dipole**  $\rightarrow NP + NP$ ; London Forces; dispersive effect

- **Ion-Induced Dipole**  $\rightarrow ion + NP$

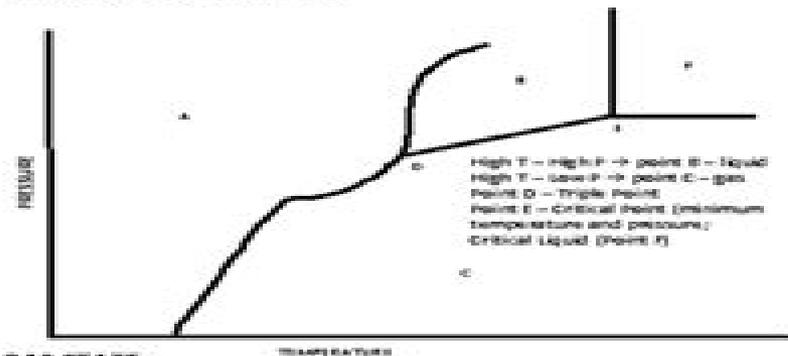
$I_2 \rightarrow NP$  (solubility = 1:2960)

$I_2 + NaI/KI$  (solubilizing agent)  $\rightarrow I_2 + I_3^-$  = increase solubility

- **Hydrogen Bond** – hydrogen atom + electronegative atom (N, O, F)



## PHASE DIAGRAM



### GAS STATE:

- **Kinetic Molecular Theory**

- Gases are composed of small particles called atoms; total volume is negligible in relation to space where they are confined
- Particles do not attract one another but move with complete independence
- Particles exhibit continuous random motion owing to their kinetic energy
- Gases exhibit perfect elasticity (no energy exchange)

## II. GAS LAWS

- at standard temperature (Kelvin =  $^{\circ}C + 273$ ) and pressure (1 atm = 760 mmHg or torr; 760 mmHg;  $1.01325 \times 10^5$  dynes/cm<sup>2</sup>;  $1.01325 \times 10^5$  Pa or N/m<sup>2</sup>; 0.1325 bar)

- **Boyle's Law** – constant temperature; increase P = decrease V; inversely related

$$P \propto \frac{1}{V}$$

- density = increase altitude; decrease volume;  $P_1V_1 = P_2V_2$

- **Charles' Law** – constant pressure; increase T = increase V

$$T \propto V; \frac{T_1}{T_2} = \frac{V_1}{V_2} \text{ or } \frac{T_1}{V_1} = \frac{T_2}{V_2}$$

- $H_2O$  – 1.0g/ml  $\rightarrow 4^{\circ}C$

- **Gay Lussac's Law**  $\rightarrow$  constant volume

$$P \propto T; \frac{P_1}{P_2} = \frac{T_1}{T_2} \text{ or } \frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- $121^{\circ}C$  – 15-20psi

# Physical Pharmacy Lecture Notes

**Michael Brown**



## **Physical Pharmacy Lecture Notes:**

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Physical Pharmacy Lecture Notes**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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