

CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron III sulfate, the roman numeral tells us **C**
 - How many atoms of Fe are in one formula unit
 - How many sulfate ions can be attached to the iron atom
 - The charge on each Fe ion
 - The total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to **c**
 - Change the number of moles represented in the formula
 - Change the charges on the other ions in the compound
 - Change the formula so that it no longer represents the compound
 - Have no effect on the formula
- Derive a generalization for when an acid name will end in the suffix **-ic** or **-ous**
If ending is -ic, the anion ended in -ate; if ending is -ous, the anion ended in -ite
- Derive a generalization for when an acid name will begin with the prefix **hydro-**.
If prefix is hydro-, the anion ended in -ide. This is most likely a monatomic ion.
- Fill in the blanks in the table below

Compound Name	Formula
Aluminum sulfide	Al_2S_3
Aluminum sulfate	$Al_2(SO_4)_3$
Lead II chloride	$PbCl_2$
Ammonium phosphite	$(NH_4)_2PO_3$
Hydroiodic acid	HCl
Dinitrogen pentoxide	N_2O_5
Copper II carbonate	$CuCO_3$
Sodium sulfide	Na_2S
Ammonium phosphate	$(NH_4)_3PO_4$
Tin IV sulfite	$Sn(SO_3)_2$
Nitrous acid	HNO_2
Magnesium perchlorate	$Mg(ClO_4)_2$
Iron II nitrate	$Fe(NO_3)_2$
Iron III nitrate	$Fe(NO_3)_3$
Cobalt II oxide	CoO
Sulfur trioxide	SO_3
Boron pentafluoride	BF_5
Tetraphosphorus octasulfide	P_4S_8
Tetraiodide nonoxide	I_4O_9

- Name the following compound and identify as molecular, ionic or acid
 - $Ca(HCO_3)_2$ **calcium bicarbonate - I**
 - $Fe(ClO_2)_3$ **Iron III chlorite - I**
 - $Fe(ClO_3)_2$ **Iron II chlorate - I**
 - N_2O_5 **dinitrogen pentoxide - M**

Chapter 9 Chemical Names Formulas Answers Key

Thomas Elliott Taylor



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Table of Contents Chapter 9 Chemical Names Formulas Answers Key

1. Understanding the eBook Chapter 9 Chemical Names Formulas Answers Key
 - The Rise of Digital Reading Chapter 9 Chemical Names Formulas Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Chemical Names Formulas Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 9 Chemical Names Formulas Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Chemical Names Formulas Answers Key
 - Personalized Recommendations
 - Chapter 9 Chemical Names Formulas Answers Key User Reviews and Ratings

- Chapter 9 Chemical Names Formulas Answers Key and Bestseller Lists
- 5. Accessing Chapter 9 Chemical Names Formulas Answers Key Free and Paid eBooks
 - Chapter 9 Chemical Names Formulas Answers Key Public Domain eBooks
 - Chapter 9 Chemical Names Formulas Answers Key eBook Subscription Services
 - Chapter 9 Chemical Names Formulas Answers Key Budget-Friendly Options
- 6. Navigating Chapter 9 Chemical Names Formulas Answers Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 9 Chemical Names Formulas Answers Key Compatibility with Devices
 - Chapter 9 Chemical Names Formulas Answers Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 9 Chemical Names Formulas Answers Key
 - Highlighting and Note-Taking Chapter 9 Chemical Names Formulas Answers Key
 - Interactive Elements Chapter 9 Chemical Names Formulas Answers Key
- 8. Staying Engaged with Chapter 9 Chemical Names Formulas Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 9 Chemical Names Formulas Answers Key
- 9. Balancing eBooks and Physical Books Chapter 9 Chemical Names Formulas Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 9 Chemical Names Formulas Answers Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 9 Chemical Names Formulas Answers Key
 - Setting Reading Goals Chapter 9 Chemical Names Formulas Answers Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 9 Chemical Names Formulas Answers Key
 - Fact-Checking eBook Content of Chapter 9 Chemical Names Formulas Answers Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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