

IB CHEMISTRY NOTES

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Atomic structure

- Electrons in order of increasing energy levels: 1s, 2s, 2p, 3s, 3p, 4s, 3d, 4p, 5s, 4d, 5p, 6s, 4f, 5d, 6p, 7s, 5f, 6d, 7p
- Electrons fill orbitals from lowest energy level (lowest n and lowest l) first
- Electrons fill orbitals from lowest energy level (lowest n and lowest l) first
- Pauli exclusion principle: no two electrons can have the same set of quantum numbers (n, l, m, s)

Energy levels

Energy levels are quantized. Electrons can only exist at higher energy levels than when it is in its ground state.

Electrons are promoted to excited states when they absorb energy. Electrons return to their ground state when they release energy.

Spectra

Electrons are promoted when photons are absorbed. Electrons return to a lower energy level when they release energy. Electrons return to their ground state when they release energy.

Atomic structure and the periodic table

Electrons fill orbitals from lowest energy level (lowest n and lowest l) first.

Energy levels and spectra

Electrons are promoted when photons are absorbed. Electrons return to a lower energy level when they release energy. Electrons return to their ground state when they release energy.

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Includes detailed handwritten notes for both SL and HL

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