

Half Modern Chemistry Review

CHAPTER 9: STOICHIOMETRY

The following pages contain the best (but not all) of the information for the chapter 9 test.

Focus on this content, but make sure to review class notes, activities, worksheets, questions, etc.

If you study this document and NOTHING else, you should at least be able to PASS the test.

**** Test items will be recall, recognition, and/or application of this content. ****

OBJECTIVES

- Calculate molar weight for molecular-formula compounds. (C 1)
- Calculate molar weight for ionic compounds. (C 1)
- 9.1. Determine moles of reactants from balanced chemical equations. (T & P)
- 9.2. Perform stoichiometry calculations with: mol to mol, mol to gram, gram to mol, and gram to gram. (P & P)
- 9.3. Calculate stoichiometry problems and determine the limiting reagent, excess reactant, and amount of product produced. (P & P)

9.1: INTRODUCTION TO STOICHIOMETRY

a. Introductory

- composition stoichiometry – calculations involving the mass relationships of elements in compounds.
- molar stoichiometry – calculations involving the mass relationships between reactants and products in a chemical reaction.
- molar ratio – a conversion factor that relates the amount, in moles, of any two substances involved in a chemical reaction.

b. Chapter Highlights

- Relates stoichiometry to the mass relationships between reactants and products in a chemical reaction.
- Stoichiometry is based on the law of conservation of mass.
- Relating mass substances to moles is called stoichiometry and expressing the amount of a given substance in moles.
- A molar ratio is the conversion factor that relates the amount in moles of any two substances in a chemical reaction. This molar ratio is derived from the balanced equation.
 - Example: $2\text{FeCl}_3 + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_4$ What is the molar ratio of FeCl_3 to Cl_2 ?
- Amount of a substance is expressed in moles, and mass of a substance is expressed using mass units such as grams, kilograms, or milligrams.
- Mass and amount of substances are quantities, whereas moles and grams are units.
- A balanced chemical equation is necessary to solve any stoichiometric problem.

9.2: MASS STOICHIOMETRIC CALCULATIONS

a. Chapter Highlights

- We use molar stoichiometric calculations, the mass, of the amount of any reactant or product over the calculation of the balanced chemical equation, and the mass or amount of any other reactant or product is known.

9.3: LIMITED REACTANTS AND PERCENTAGE YIELD

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Yicheng Fang



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