

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the the systems, which are exploring the structures through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}^{+}(\text{aq}) \rightleftharpoons \text{FeSCNH}^{3+}(\text{aq})$$
4. In part 4, what is the purpose of adding potassium thiocyanate to the solution of cobalt ions in water?
Cobalt ions form an octahedral complex and have a color change but nothing. The purpose of adding cobalt is to compare the color of the solution and determine what color the cobalt hexahydrate is.
5. What is the purpose of adding distilled water in part 5?
Water is added to dilute the solution of the cobalt to make subsequent changes easier to see.
6. What is the purpose of the test tube in part 6?
To observe the color change of the cobalt complex as the temperature changes.
7. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$$
8. Write any reactions that Fe^{3+} can undergo with OH^{-} ions.
Cobalt complex Co^{2+} will react the hydroxide ions to produce cobalt hydroxide $\text{Co}(\text{OH})_2$ and cobalt ions Fe^{3+} will produce $\text{Fe}(\text{OH})_3$.
9. What is a precipitate?
A precipitate is a solid that forms from a solution. It is a solid that is not dissolved in the solution. It is a solid that is not dissolved in the solution. It is a solid that is not dissolved in the solution.

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L Darling-Hammond

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