

SEA CADETS

Navy League Cadets (Age 10-12)

Sea Cadets (Age 12-18)

Sea Cadet Officers (Age 12-18)

<p>LC-7 Ship's Leading Petty Officer (SLPO)</p> 	<p>SC-7 Chief Petty Officer (CPO)</p> 	<p>O-4 Lieutenant Commander (LCDR)</p>  
<p>LC-6 Petty Officer First Class (PO1)</p> 	<p>SC-6 Petty Officer First Class (PO1)</p> 	<p>O-3 Lieutenant (LT)</p>  
<p>LC-5 Petty Officer Second Class (PO2)</p> 	<p>SC-5 Petty Officer Second Class (PO2)</p> 	<p>O-2 Lieutenant Junior Grade (LTJG)</p>  
<p>LC-4 Petty Officer Third Class (PO3)</p> 	<p>SC-4 Petty Officer Third Class (PO3)</p> 	<p>O-1 Ensign (ENS)</p>  
<p>LC-3 Able Cadet</p> 	<p>SC-3 Seaman (SN)</p> 	<p>W-2 Warrant Officer (CWO)</p>  
<p>LC-2 Apprentice Cadet</p> 	<p>SC-2 Seaman Apprentice (SA)</p> 	<p>--- Midshipman (MIDN)</p>  
<p>LC-2 Recruit Cadet</p> <p>No Insignia</p>	<p>SC-1 Seaman Recruit (SR)</p> <p>No Insignia</p>	<p>--- Instructor (INST)</p>  <p>No Sleeve or Shoulder Insignia</p>

Sea Cadet Seaman Course Work Answer Key

S Marginson



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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the

number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass
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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to
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 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$ one mole of HCl would be fully neutralized by one mole of NaOH

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction naoh aq c6h4 cooh cook aq c6h4 coona cook aq h2o l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the

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