

## Pre-lab: Acid-Base Titrations

CHM205 Laboratory

Fall 2019 Section: \_\_\_\_\_

Grade: \_\_\_\_\_

Name: \_\_\_\_\_

A procedure similar to the one in this lab is used to titrate a sodium hydroxide solution against potassium hydrogen phthalate (KHP). The follow data is recorded: For the standard KHP solution, 10.575 g of KHP is dissolved in water, and diluted to 250.00 mL in a volumetric flask; 20.00 mL of the KHP solution are added to an Erlenmeyer flask; a buret is filled with a NaOH solution of unknown concentration to a starting volume of 0.18 mL, and at the end point of the titration the buret reading is 20.83 mL.

1. Calculate the mmoles KHP that were put into the 250.00 mL volumetric flask. Note: KHP's chemical formula is  $\text{KHC}_8\text{H}_4\text{O}_4$ .
2. Calculate the KHP solution concentration (in molarity units) that was prepared in the 250 mL volumetric flask. Note sig figs on the flask volume.
3. Calculate the mmoles KHP that are in the 20.00 mL aliquot.
4. Write the balanced chemical equation for the reaction of KHP with sodium hydroxide.
5. Calculate the mmoles NaOH that will be needed to reach the titration end point, when the 20.00 mL aliquot is completely titrated.
6. Calculate the volume of NaOH used in the titration.
7. Calculate the concentration (in molarity units) of the NaOH solution.

# Acid Base Titrations Pre Lab Answers

**Clemens Wendtner**



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