

## CHAPTER - 2

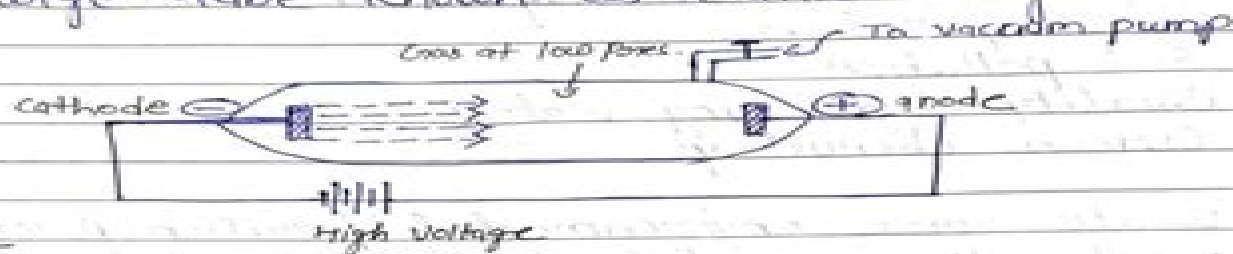
### STRUCTURE OF ATOM

#### Discovery of electron (Cathode Rays)

Cathode rays  $\rightarrow$  William Crookes (1879)

Electron ( $e^-$ )  $\rightarrow$  J. J. Thomson (-ve charge particle)

-ve charged rays which move from cathode to anode in discharge tube known as cathode rays.



#### properties

- 1 Cathode rays travel in a straight line (shadow)
- 2 Cathode rays carry -ve charge (electrostatic field deflection)
- 3 Cathode rays are deflected by magnetic field (rotate paddle wheel)
- 4 They are made up of material particle
- 5 They produce heating effect & possess penetrating effect
- 6 They produce green fluorescence & X-rays (when they strike)
- 7 They also affect photographic plate.

#### Discovery of proton (Anode / Canal Rays)

Anode rays  $\rightarrow$  Goldstein (1886)

Proton ( $p^+$ )  $\rightarrow$  Goldstein & named by Rutherford (+ve charge particle)

+ve charged particle which move opposite the cathode in discharge tube known as anode rays.

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**Richard Bailey**



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magnificat nunc dimittis responses anthem hymn s organ voluntary  
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web view order of service monday 11th september 2023 5 00pm evensong quire sung by the choristers plainsong lucis creator optime plainsong responses moore second service haydn benedictus missa brevis sancti joannis de deo view order of service tuesday 12th september 2023 5 00pm

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