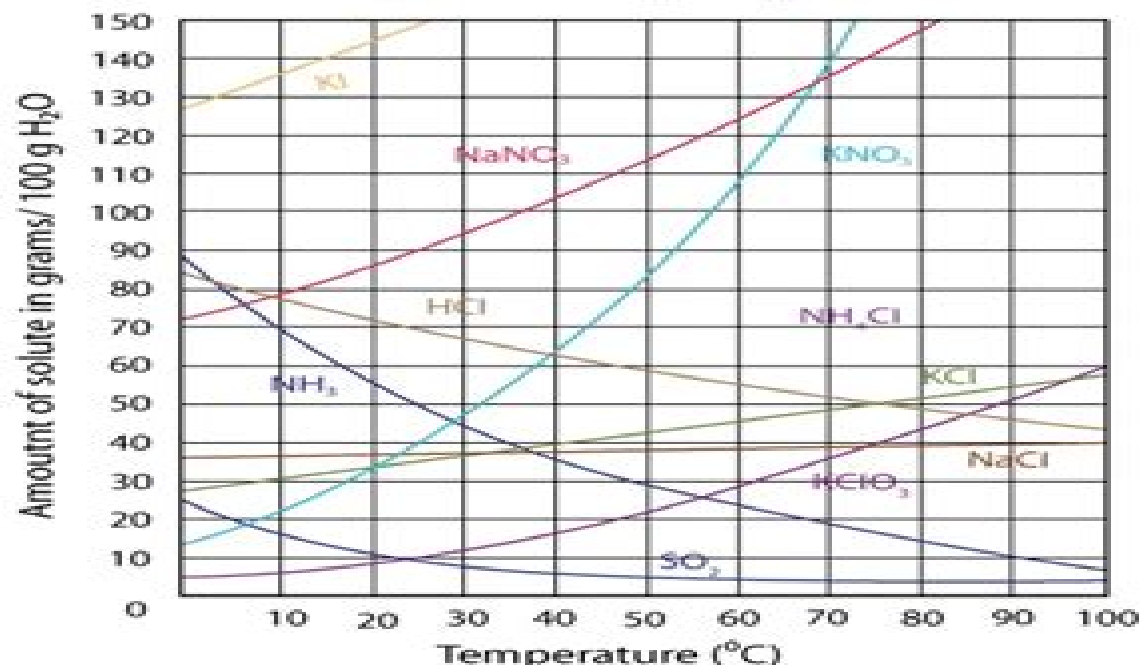




Interpreting Solubility Curve



1. Define solubility. _____
2. Using the solubility curve below, answer the given questions.



- I. What is the solvent in this solubility curve? _____
- II. Saturation points are given per how much of the solvent? _____
- III. What is the saturation point of NaNO₃ at 10°C, 40°C, & 80°C?
 10°C = _____ 40°C = _____ 80°C = _____
- IV. How much NH₃ can you dissolve in water at 10°C, 30°C, & 90°C?
 10°C = _____ 30°C = _____ 90°C = _____
- V. Which salt is the least soluble in water? _____
- VI. How many grams of potassium chloride can be dissolved in 200 g of water at 80 °C? _____
- VII. At 40 °C, how much potassium nitrate will be dissolved in 300 g of water? _____
- VIII. Which salt shows the least change in solubility from 0° to 100° C? _____
- IX. Which substance shows a decrease in solubility from 0° to 100° C? _____
- X. At 30 °C, 90 g of sodium nitrate is dissolved in 100 g water. Is this solution saturated, unsaturated, or supersaturated? _____

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web 1 a 4.31×10^{22} atoms ni b 7.01×10^{20} ions br c 2.52×10^{23} molecules h 2 d 7.79×10^{21} molecules h 2 o 2 2 a 5 25 g naoh b 1 15 x 10 3 g lactose c 82 1 g fe d 0 332 g no 2 3 6 54 x 10 21 molecules of c 3 h 8 o 3 4 2 79 grams of nh 3 5 1 97 x 10 22 grams of sn

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