

## CHAPTER

## STUDY GUIDE

## Covalent Bonding

### Section 8.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond

molecule

sigma bond

exothermic

pi bond

When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1)** \_\_\_\_\_. When such an attachment is formed, bond dissociation energy is released, and the process is **(2)** \_\_\_\_\_. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3)** \_\_\_\_\_. If the electrons shared are centered between the two atoms, the attachment is called a(n) **(4)** \_\_\_\_\_. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5)** \_\_\_\_\_.

In your textbook, read about single and multiple bonds and bond strength.

Circle the letter of the choice that best completes the statement or answers the question.

6. In what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
- as single atoms
  - as molecules containing two atoms
  - as molecules containing three atoms
  - as molecules containing four atoms
7. How many electrons are shared in a double covalent bond?
- none
  - one
  - two
  - four
8. Bond length is the distance between
- two molecules of the same substance.
  - the electrons in two attached atoms.
  - the nuclei of two attached atoms.
  - the orbitals of two attached atoms.
9. Which of the following relationships relating to bond length is generally correct?
- the shorter the bond, the stronger the bond
  - the shorter the bond, the weaker the bond
  - the shorter the bond, the fewer the electrons in it
  - the shorter the bond, the lower the bond dissociation energy

# Chapter 8 Covalent Bonding Study Guide Answers

**E Durkheim**



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## **Table of Contents Chapter 8 Covalent Bonding Study Guide Answers**

1. Understanding the eBook Chapter 8 Covalent Bonding Study Guide Answers
  - The Rise of Digital Reading Chapter 8 Covalent Bonding Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 8 Covalent Bonding Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Study Guide Answers
  - Personalized Recommendations
  - Chapter 8 Covalent Bonding Study Guide Answers User Reviews and Ratings
  - Chapter 8 Covalent Bonding Study Guide Answers and Bestseller Lists
5. Accessing Chapter 8 Covalent Bonding Study Guide Answers Free and Paid eBooks
  - Chapter 8 Covalent Bonding Study Guide Answers Public Domain eBooks
  - Chapter 8 Covalent Bonding Study Guide Answers eBook Subscription Services
  - Chapter 8 Covalent Bonding Study Guide Answers Budget-Friendly Options

6. Navigating Chapter 8 Covalent Bonding Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 8 Covalent Bonding Study Guide Answers Compatibility with Devices
  - Chapter 8 Covalent Bonding Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Study Guide Answers
  - Highlighting and Note-Taking Chapter 8 Covalent Bonding Study Guide Answers
  - Interactive Elements Chapter 8 Covalent Bonding Study Guide Answers
8. Staying Engaged with Chapter 8 Covalent Bonding Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 8 Covalent Bonding Study Guide Answers
9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Study Guide Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Study Guide Answers
  - Setting Reading Goals Chapter 8 Covalent Bonding Study Guide Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Study Guide Answers
  - Fact-Checking eBook Content of Chapter 8 Covalent Bonding Study Guide Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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