

## Chapter 9 Practice Problems: Chemical Names and Formulas

### Section 9.1: Naming Ions

- What is the charge on the ion typically formed by each element?
  - oxygen
  - iodine
  - sodium
  - aluminum
  - nickel, two electrons lost
  - magnesium
- How many electrons does the neutral atom gain or lose when each ion forms?
  - $\text{Cr}^{3+}$
  - $\text{P}^{3-}$
  - $\text{Li}^+$
  - $\text{Ca}^{2+}$
  - $\text{Cl}^-$
  - $\text{O}^{2-}$
- Name each ion. Identify each as a cation or an anion.
  - $\text{Sn}^{2+}$
  - $\text{Co}^{3+}$
  - $\text{Br}^-$
  - $\text{K}^+$
  - $\text{H}^-$
  - $\text{Mn}^{2+}$
- Write the formula (including the charge) for each ion. Use table 9.3 as if necessary.
  - carbonate ion
  - nitrite ion
  - sulfate ion
  - hydroxide ion
  - chromate ion
  - ammonium ion
- Name each of the following ions. Identify each as a cation or an anion.
  - $\text{CN}^-$
  - $\text{HCO}_3^-$
  - $\text{PO}_4^{3-}$
  - $\text{Cl}^-$
  - $\text{Ca}^{2+}$
  - $\text{SO}_3^{2-}$

### 9.2 Naming and Writing Formulas for Ionic Compounds:

- Write the formulas for these binary ionic compounds.
  - magnesium oxide
  - tin(II) fluoride
  - potassium iodide
  - aluminum chloride
  - sodium sulfide
  - ferric bromide
- Write the formulas for the compounds formed from these pairs of ions.
  - $\text{Ba}^{2+}$  and  $\text{Cl}^-$
  - $\text{Ag}^+$  and  $\text{I}^-$
  - $\text{Ca}^{2+}$  and  $\text{S}^{2-}$
  - $\text{K}^+$  and  $\text{Br}^-$
  - $\text{Al}^{3+}$  and  $\text{O}^{2-}$
  - $\text{Fe}^{2+}$  and  $\text{O}^{2-}$
- Name the following binary ionic compounds.
  - $\text{MnO}_2$
  - $\text{Li}_3\text{N}$
  - $\text{CaCl}_2$
  - $\text{SrBr}_2$
  - $\text{NiCl}_2$
  - $\text{K}_2\text{S}$
  - $\text{Fe}_3\text{N}_2$
  - $\text{SnCl}_4$
- Write formulas for the following ionic compounds.
  - sodium phosphate
  - magnesium sulfate
  - sodium hydroxide
  - potassium cyanide
  - ammonium chloride
  - potassium dichromate
- Write formulas for the following ionic compounds.
  - $\text{NH}_4^+$  and  $\text{SO}_4^{2-}$
  - $\text{K}^+$  and  $\text{NO}_3^-$
  - barium ion and hydroxide ion
  - lithium ion and carbonate ion

# Chapter 9 Chemical Names Formulas Practice Problems Answers

**SB Merriam**



## Chapter 9 Chemical Names Formulas Practice Problems Answers:

*General Organic and Biological Chemistry* Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features [Student's Guide to Introduction to Chemical Principles](#) by Edward I. Peters, 2d Ed Peter P. Berlow, 1978 **Understanding Chemistry** Fred M. Dewey, 1994 [Fundamentals Davis Essential Nursing Content + Practice Questions](#) Patricia M Nugent, Barbara A Vitale, 2017-04-20 Up to date with the NCLEX RN 2016 Test Plan Too much information Too little time Here s everything you need to succeed in your fundamentals of nursing course and prepare for course exams and the NCLEX Succinct content review in outline format focus on must know information while case studies and NCLEX style questions develop your ability to apply your knowledge in simulated clinical situations You ll also find proven techniques and tips to help you study more effectively learn how to approach different types of questions and improve your critical thinking skills **eBook: General, Organic and Biological Chemistry 2e** SMITH, 2012-02-16 eBook *General Organic and Biological Chemistry 2e* [Merrill Chemistry](#) Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Platinotype** Pradip Malde, Mike Ware, 2020-12-29 *Platinotype Making Photographs in Platinum and Palladium with the Contemporary Printing out Process* describes the mechanisms and chemistry of platinum palladium printing in safe and practical ways Clearly presented formulae allow the printer to work with platinum palladium or varying combinations of both The printed out image appears fully during exposure and only requires simple and safe steps for clearing to a stable archival state The authors explain what makes the image how all necessary components are prepared and used and the kind of paper and negative needed to make prints More than just a technical manual the book underscores the authors belief that printing is a creative scientific and philosophic way of working The book presents an outstanding collection of prints by over 40 artists all made with this printing out process The artists notes and comments offer insights into their methods and thinking and a large number of full page reproductions serve as a valuable reference for the aspiring printer The book includes A list of supplies and equipment A detailed chemical glossary A Quick start section in the Preface Summary sheets and workflows for each step of the process Instructions for making traditional negatives with Pyro PMK and digital negatives Explanation of the chemistry and dynamics of paper and how to use buffered papers Instructions for controlling hydration processes and humidity Instructions for preparing each chemical solution needed for the process Discussion about the aesthetics of the platinum palladium print Explanation of the relationships between light image and expression A detailed troubleshooting list Recommendations from conservators about processing handling and conservation Contemporary artists using the printing out platinum palladium process Learning how

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