

7-2 Review and Reinforcement

Covalent Bonding

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line provided.

 F, covalent 1. A group of atoms united by ionic bonds is called a molecule.

 T 2. A covalent bond is formed by a shared pair of electrons

 F, four 3. A double covalent bond consists of two shared electrons.

 T 4. A molecular formula tells you how many atoms are in a single molecule of the compound.

 F, MF 5. The empirical formula for a molecule specifies which atoms are bonded to each other in the molecule.

 F, lone pair 6. A pair of electrons not involved in bonding is called a shared electron pair.

 F, equally 7. In nonpolar covalent bonds, the electrons are shared unequally between two atoms.

 F, Triple 8. Lewis structures use a triple dash to represent a double bond.

Use the information in Figure 7-19 of the textbook to calculate the electronegativity differences between each pair of elements listed in the chart below. Then indicate whether the bond would be nonpolar covalent, polar covalent, or ionic. Write your answers in the chart.

		Electronegativity Differences	Type of Bond
9.	H-O		
10.	C-H		
11.	K-F		
12.	N-H		
13.	Na-F		
14.	O-Cl		

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**Thirumalaisamy P.
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Kid's Box American English Level 3 Teacher's Resource Pack with Audio CD Kathryn Escribano, Caroline

Nixon, 2011-03-10 Kid's Box is a six level course for young learners Bursting with bright ideas to inspire both teachers and students Kid's Box American English gives children a confident start to learning English It also fully covers the syllabus for the Cambridge Young Learners English YLE tests This Resource Pack contains extra activities to reinforce and extend each unit of the Student's Book allowing teachers to cater for mixed ability classes as well as tests suitable for YLE preparation It is accompanied by an Audio CD complete with songs listening exercises and tests Level 3 begins the Movers cycle CEF level

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