

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
 potassium-40 undergoes beta decay (electron)  
 uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.  
 Write the electron configuration for Cr.  
 Write the shorthand notation for Bi.



### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds:

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



# Modern Chemistry Chapter 7 Section 1 Review Answers

**Norma K Dunlap, Donna M Huryn**



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*Medicinal Chemistry* Norma K Dunlap, Donna M Huryn, 2025-11-28 This book provides a concise introduction to the medicinal chemistry of therapeutically active compounds Written in an accessible style it offers a brief overview of the drug development process from discovery to approved drug Emphasizing medicinal chemistry strategies the content is organized by target as well as therapeutic area drug classes It covers existing drugs and the discovery process that led to their development Aimed at students and professionals interested in drug discovery Medicinal Chemistry requires little prior knowledge making it ideal for classroom and training seminar use Written by experts in the field this outstanding new edition explains medicinal chemistry on chemical mechanism terms allowing for a more interesting and user friendly text Includes valuable insights toward the various pathways taken by the pharmaceutical industry in drug discoveries Enhanced text includes questions raised and suggestions made from students in the authors medicinal chemistry class Content provided to enhance the text includes a solutions manual PowerPoint slides for lecture outlines and especially for the earlier chapters a selection of videos An accompanying website will be updated to include additional material 1 Journal Club exercises current journal articles with guiding questions for student analysis two per chapter 2 additional case studies 3 organic synthesis analysis selected syntheses both discovery and process will be discussed and 4 updates on current advances in the field of medicinal chemistry The book benefits a diverse audience including upper level undergraduates and graduates who are studying in the fields of medicinal chemistry and drug discovery as well as scientists working in the pharmaceutical industry

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