

Stoichiometry: Mass – Mole Worksheet

Solve the following stoichiometry problems and type in the correct answer with units.
(for example: 3.45 mol or 4.62 g) Round the masses from the periodic table to the nearest whole number. Round all answers to 2 decimal places.

1. Based on the following chemical equation, please answer the following questions



- a. Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₂.

Answer: _____ $\frac{\text{g FeCl}_2}{1 \text{ mol FeCl}_2} \times \frac{1 \text{ mol FeCl}_2}{\text{g FeCl}_2} \times \frac{\text{mol NaCl}}{\text{mol FeCl}_2} =$

- b. Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

Answer: _____ $\frac{\text{g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \times \frac{1 \text{ mol Na}_2\text{SO}_4}{\text{g Na}_2\text{SO}_4} \times \frac{\text{mol Fe}_2(\text{SO}_4)_3}{\text{mol Na}_2\text{SO}_4} \times \frac{\text{g Fe}_2(\text{SO}_4)_3}{1 \text{ mol Fe}_2(\text{SO}_4)_3} =$

- c. Calculate the amount in grams of FeCl₂ produced with 95g of Fe(SO₄)₂.

Answer: _____ $\frac{\text{g Fe}(\text{SO}_4)_2}{1 \text{ mol Fe}(\text{SO}_4)_2} \times \frac{1 \text{ mol Fe}(\text{SO}_4)_2}{\text{g Fe}(\text{SO}_4)_2} \times \frac{\text{mol FeCl}_2}{\text{mol Fe}(\text{SO}_4)_2} \times \frac{\text{g FeCl}_2}{1 \text{ mol FeCl}_2} =$

- d. Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl.

Answer: _____ $\frac{\text{g NaCl}}{1 \text{ mol NaCl}} \times \frac{1 \text{ mol NaCl}}{\text{g NaCl}} \times \frac{\text{mol Fe}_2(\text{SO}_4)_3}{\text{mol NaCl}} =$

2. How many moles of CaCO₃ can be produced if we make 12.3 g of Ca(HCO₃)₂ reacts with CaCO₃?



Answer: _____ $\frac{\text{g Ca}(\text{HCO}_3)_2}{1 \text{ mol Ca}(\text{HCO}_3)_2} \times \frac{1 \text{ mol Ca}(\text{HCO}_3)_2}{\text{g Ca}(\text{HCO}_3)_2} \times \frac{\text{mol CaCO}_3}{\text{mol Ca}(\text{HCO}_3)_2} =$

Moles And Stoichiometry Packet Answers

DJ Losen



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