

## Assessment

## Chapter Test A

## Chapter: Acids and Bases

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. Which of the properties listed below is *not* characteristic of an acid?
- a sour taste
  - a slippery feel
  - the ability to conduct an electric current
  - reactivity with metals
- \_\_\_\_\_ 2. Which is an example of a weak acid?
- HCl
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_3\text{PO}_4$
  - $\text{HNO}_3$
- \_\_\_\_\_ 3. Acetic acid is found in
- vinegar.
  - the stomach.
  - antacids.
  - oranges.
- \_\_\_\_\_ 4. A characteristic of an Arrhenius base is that it
- is an electrolyte.
  - tastes bitter.
  - dissociates to form  $\text{OH}^-$  ions.
  - All of the above
- \_\_\_\_\_ 5. Which of the following is a Brønsted-Lowry base?
- HCl
  - $\text{HCO}_3^-$
  - $\text{H}_3\text{O}^+$
  - $\text{H}_3\text{PO}_4$
- \_\_\_\_\_ 6. In the reaction represented by the equation
- $$\text{H}_2\text{C}_2\text{O}_4(\text{aq}) + \text{CH}_3\text{NH}_2(\text{aq}) \rightleftharpoons \text{HC}_2\text{O}_4^-(\text{aq}) + \text{CH}_3\text{NH}_3^+(\text{aq}),$$
- which of these is a conjugate acid-base pair?
- $\text{H}_2\text{C}_2\text{O}_4(\text{aq})$  and  $\text{CH}_3\text{NH}_2(\text{aq})$
  - $\text{H}_2\text{C}_2\text{O}_4(\text{aq})$  and  $\text{CH}_3\text{NH}_3^+(\text{aq})$
  - $\text{CH}_3\text{NH}_2(\text{aq})$  and  $\text{CH}_3\text{NH}_3^+(\text{aq})$
  - $\text{HC}_2\text{O}_4^-(\text{aq})$  and  $\text{CH}_3\text{NH}_3^+(\text{aq})$

# Modern Chemistry Chapter 7 Test Answer Key

**Tobias Bleicker**



## **Modern Chemistry Chapter 7 Test Answer Key:**

Elementary Modern Chemistry Wilhelm Ostwald, Harry Wheeler Morse, 1909      **Modern Electrochemistry 2B** John O'M. Bockris, Amulya K.N. Reddy, Maria E. Gamboa-Aldeco, 1998 This long awaited and thoroughly updated version of the classic text Plenum Press 1970 explains the subject of electrochemistry in clear straightforward language for undergraduates and mature scientists who want to understand solutions Like its predecessor the new text presents the electrochemistry of solutions at the molecular level The Second Edition takes full advantage of the advances in microscopy computing power and industrial applications in the quarter century since the publication of the First Edition Such new techniques include scanning tunneling microscopy which enables us to see atoms on electrodes and new computers capable of molecular dynamics calculations that are used in arriving at experimental values Chapter 10 starts with a detailed description of what happens when light strikes semi conductor electrodes and splits water thus providing in hydrogen a clean fuel There have of course been revolutionary advances here since the First Edition was written The book also discusses electrochemical methods that may provide the most economical path to many new syntheses for example the synthesis of the textile nylon The broad area of the breakdown of material in moist air and its electrochemistry is taken up in the substantial Chapter 12 Another exciting topic covered is the evolution of energy conversion and storage which lie at the cutting edge of clean automobile development Chapter 14 presents from a fresh perspective a discussion of electrochemical mechanisms in Biology and Chapter 15 shows how new electrochemical approaches may potentially alleviate many environmental problems      **Lab Experiments Modern Chemistry** Harold Clark Metcalfe, 1982      **Modern Inorganic Chemistry** Joseph William Mellor, 1917      Introduction to Corrosion Science Edward McCafferty, 2010-07-17 This textbook is intended for a one semester course in corrosion science at the graduate or advanced undergraduate level The approach is that of a physical chemist or materials scientist and the text is geared toward students of chemistry materials science and engineering This textbook should also be useful to practicing corrosion engineers or materials engineers who wish to enhance their understanding of the fundamental principles of corrosion science It is assumed that the student or reader does not have a background in electrochemistry However the student or reader should have taken at least an undergraduate course in materials science or physical chemistry More material is presented in the textbook than can be covered in a one semester course so the book is intended for both the classroom and as a source book for further use This book grew out of classroom lectures which the author presented between 1982 and the present while a professorial lecturer at George Washington University Washington DC where he organized and taught a graduate course on Environmental Effects on Materials Additional material has been provided by over 30 years of experience in corrosion research largely at the Naval Research Laboratory Washington DC and also at the Bethlehem Steel Company Bethlehem PA and as a Robert A Welch Postdoctoral Fellow at the University of Texas The text emphasizes basic principles of corrosion science which underpin

extensions to practice     Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance     *Laboratory Experiments to Accompany "Modern Chemistry,"* Fredus Nelson Peters, 1902     **Practical Chemistry** Newton Henry Black, James Bryant Conant, 1920     **Books in Print Supplement**, 2002     Modern Chemistry Nicholas D. Tzimopoulos, 1990     **Handbook of Modern Chemistry** Charles Meymott Tidy, 1878     **Modern Microcrystal Tests for Drugs** Charles Clarke Fulton, 1969     **General Chemistry** Ralph H. Petrucci, 1989     **Lab Experiments for Modern Chemistry** Tzimopoulo, 1990     **Children's Books in Print** R R Bowker Publishing, Bowker, 1999-12     **Introduction to Modern Chemistry: Student Manual** Eugene Meyer, 1979     **Modern Chemistry** H. Clark Metcalfe, Joseph F. Castka, 1966     Essentials of Modern Chemistry Charles Elwood Dull, 1918     **Modern Inorganic Chemistry** William L. Jolly, 1984     **Handbook of Modern Pharmaceutical Analysis** Satinder Ahuja, Stephen Scypinski, 2010-11-11 Handbook of Modern Pharmaceutical Analysis Second Edition synthesizes the complex research and recent changes in the field while covering the techniques and technology required for today's laboratories The work integrates strategy case studies methodologies and implications of new regulatory structures providing complete coverage of quality assurance from the point of discovery to the point of use Treats pharmaceutical analysis PA as an integral partner to the drug development process rather than as a service to it Covers method development validation selection testing modeling and simulation studies combined with advanced exploration of assays impurity testing biomolecules and chiral separations Features detailed coverage of QA ethics and regulatory guidance quality by design good manufacturing practice as well as high tech methodologies and technologies from lab on a chip to LC MS LC NMR and LC NMR MS

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