

## CHAPTER 8 REVIEW

# Chemical Equations and Reactions

**SHORT ANSWER** Answer the following questions in the space provided.

2. \_\_\_\_\_ According to the law of conservation of mass, the total mass of the reacting substances is
- always more than the total mass of the products.
  - always less than the total mass of the products.
  - sometimes more and sometimes less than the total mass of the products.
  - always equal to the total mass of the products.**

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 2. In the reaction described by the word equation  

$$\text{sodium} + \text{water} \rightarrow \text{sodium hydroxide} + \text{hydrogen gas}$$
  - sodium hydroxide is a product.
  - sodium is a product.
  - water is a reactant.
  - Both (a) and (c).**
- \_\_\_\_\_ 3. Which of the following indicates that a chemical equation is balanced?  
  - The numbers of atoms of each element are the same on both sides of the equation.**
  - All of the coefficients are the same.
  - The numbers of molecules on each side of the equation are equal.
  - The sums of the coefficients on each side of the equation are equal.
- \_\_\_\_\_ 4. From a complete and correctly written chemical equation, you can obtain the  
  - chemical formulas of the reactants and products.
  - relative amounts of the reactants and products.
  - physical states of the reactants and products.
  - All of the above.**
- \_\_\_\_\_ 5. To balance a chemical equation, you must  
  - adjust the subscripts.
  - adjust the number of products.
  - adjust the number of reactants.
  - adjust the coefficients.**
- \_\_\_\_\_ 6. The formation of a solid product in a chemical reaction is represented by the symbol  
  - (aq).
  - (g).
  - (l).
  - (s).**
- \_\_\_\_\_ 8. In the expression  $3\text{CO}_2$ , the numbers 3 and 2 are, respectively,  
  - a subscript and a coefficient.
  - a coefficient and a subscript.**
  - two subscripts.
  - two coefficients.
- \_\_\_\_\_ 9. What is the total number of molecules that is represented by the following chemical equation? (Add up the coefficients.)  

$$\text{NH}_4\text{NO}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$$
  - 3
  - 4**
  - 8
  - 16
- \_\_\_\_\_ 10. Which equation below violates the law of conservation of mass?  
  - $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
  - $\text{KCl} + \text{Br} \rightarrow \text{KBr} + \text{Cl}_2$**
  - $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$
  - $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$
- \_\_\_\_\_ 11. Which of the following equations represents the balanced equation for the reaction of iron and oxygen?  
  - $2\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
  - $\text{Fe} + 3\text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
  - $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$**
  - $3\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
- \_\_\_\_\_ 12. Which coefficients correctly balance the formula equation below?  

$$\text{KClO}_3(\text{s}) \rightarrow \text{KCl}(\text{s}) + \text{O}_2(\text{g})$$
  - 1, 1, 1
  - 1, 1, 3
  - 2, 2, 3**
  - 2, 1, 1
- \_\_\_\_\_ 13. Which equation is *not* balanced?  
  - $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
  - $4\text{H}_2 + 2\text{O}_2 \rightarrow 4\text{H}_2\text{O}$
  - $\text{H}_2 + \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$
  - $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$**
- \_\_\_\_\_ 14. In which type of reaction do two or more compounds react to form one product?  
  - synthesis (AKA combination)**
  - decomposition
  - combustion
  - single-replacement
- \_\_\_\_\_ 15. Which of the following is true regarding single-replacement reactions?  
  - Only one product is formed.
  - There is only one reagent.
  - It requires oxygen.
  - One element replaces another in a compound.**

# Chapter 8 Chemistry Answers

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