

Chapter 17: Chemical Equilibrium

Worksheet: K_c and K_p

1. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$

at 1000 K

- 1. The equilibrium constant K_c for the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$ is 16.0 at 1000 K. Calculate K_c for the reaction $\text{SO}_2(g) + \frac{1}{2}\text{O}_2(g) \rightleftharpoons \text{SO}_3(g)$ at 1000 K.
- 2. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, $K_c = 16.0$ at 1000 K. Calculate K_c for the reaction $\text{SO}_3(g) \rightleftharpoons \text{SO}_2(g) + \frac{1}{2}\text{O}_2(g)$ at 1000 K.
- 3. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, $K_c = 16.0$ at 1000 K. Calculate K_c for the reaction $\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons \text{SO}_3(g)$ at 1000 K.
- 4. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, $K_c = 16.0$ at 1000 K. Calculate K_c for the reaction $\text{SO}_3(g) \rightleftharpoons \text{SO}_2(g) + \text{O}_2(g)$ at 1000 K.
- 5. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, $K_c = 16.0$ at 1000 K. Calculate K_c for the reaction $\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons \text{SO}_3(g)$ at 1000 K.
- 6. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, $K_c = 16.0$ at 1000 K. Calculate K_c for the reaction $\text{SO}_3(g) \rightleftharpoons \text{SO}_2(g) + \frac{1}{2}\text{O}_2(g)$ at 1000 K.
- 7. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, $K_c = 16.0$ at 1000 K. Calculate K_c for the reaction $\text{SO}_2(g) + \frac{1}{2}\text{O}_2(g) \rightleftharpoons \text{SO}_3(g)$ at 1000 K.
- 8. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, $K_c = 16.0$ at 1000 K. Calculate K_c for the reaction $\text{SO}_3(g) \rightleftharpoons \text{SO}_2(g) + \text{O}_2(g)$ at 1000 K.
- 9. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, $K_c = 16.0$ at 1000 K. Calculate K_c for the reaction $\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons \text{SO}_3(g)$ at 1000 K.
- 10. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, $K_c = 16.0$ at 1000 K. Calculate K_c for the reaction $\text{SO}_3(g) \rightleftharpoons \text{SO}_2(g) + \frac{1}{2}\text{O}_2(g)$ at 1000 K.

Answers

- 1. $K_c = 0.25$
- 2. $K_c = 0.0625$
- 3. $K_c = 0.25$
- 4. $K_c = 0.0625$
- 5. $K_c = 0.25$
- 6. $K_c = 0.0625$
- 7. $K_c = 0.25$
- 8. $K_c = 0.0625$
- 9. $K_c = 0.25$
- 10. $K_c = 0.0625$

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Katrin Zwirgmaier



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