

CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1** - Atomic Number
- H** - Atomic Symbol
- 1.008** - Atomic Weight



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For neutral atoms, $Z = +$ the number of electrons. The mass number, A , is the sum of the number of protons and neutrons. Isotopes of an element have the same Z , different A . The atomic mass is the weighted average of isotopes. See also: Atomic Mass, Atomic Weight, Crystal Structure, Periodic Table of Elements.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal Quantum Number (n): $n = 1, 2, 3, \dots$
Angular Momentum Quantum Number (l): $l = 0, 1, 2, \dots, n-1$
Magnetic Quantum Number (m_l): $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$
Spin Quantum Number (m_s): $m_s = +\frac{1}{2}, -\frac{1}{2}$

NUCLEAR CHEMISTRY

Nuclear reactions alter the number, particles, energy, and mass of atoms. Nuclear reactions are studied in nuclear physics and nuclear chemistry. Energy (E) is related to mass (m) by the equation $E = mc^2$. Nuclear reactions are studied in nuclear physics and nuclear chemistry. Energy (E) is related to mass (m) by the equation $E = mc^2$. Nuclear reactions are studied in nuclear physics and nuclear chemistry.

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the behavior of electrons in many-electron atoms and ions. A many-electron atom has many different energy levels and orbital properties. The main point is that orbitals have different energies. The filling of levels is guided by the Aufbau Principle.

n	l	m_l	m_s	Orbital
1	0	0	$\pm \frac{1}{2}$	1s orbital
2	0	0	$\pm \frac{1}{2}$	2s orbital
2	1	0	$\pm \frac{1}{2}$	2p _z orbital
2	1	± 1	$\pm \frac{1}{2}$	2p _x and 2p _y orbitals
3	0	0	$\pm \frac{1}{2}$	3s orbital
3	1	0	$\pm \frac{1}{2}$	3p _z orbital
3	1	± 1	$\pm \frac{1}{2}$	3p _x and 3p _y orbitals
3	2	0	$\pm \frac{1}{2}$	3d _{z²} orbital
3	2	± 1	$\pm \frac{1}{2}$	3d _{xy} and 3d _{xz} orbitals
3	2	± 2	$\pm \frac{1}{2}$	3d _{yz} and 3d _{x²-y²} orbitals

Atomic Configuration: Shows orbital occupancy. Pauli Exclusion Principle: Each orbital has a unique set of quantum numbers. Aufbau Principle: Orbitals are filled in order of increasing energy. Hund's Rule: Electrons occupy orbitals of equal energy singly before pairing up.



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