

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation:	<u> 2 </u> H ₂ S _(g)	+	<u> 3 </u> O _{2(g)}	→	<u> 2 </u> SO _{2(g)}	+	<u> 2 </u> H ₂ O _(g)
Before	<u> 2 </u>		<u> ex </u>		<u> 0 </u>		<u> 0 </u> 3.0 mol O ₂
Change	<u> -2 </u>		<u> -3 </u>		<u> 2 </u>		<u> 2 </u>
After	<u> 0 </u>		<u> ex </u>		<u> 2 </u>		<u> 2 </u>

- Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$

Before	0.5	ex	0	0	1.50mol HCL
Change	-0.5	-1.5	+0.5	+0.5	
After	0	ex	0.5	0.75	

- Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before	ex	1.25	0	2.50 mole
Change	-3.75	-1.25	+25	
After	ex	0	2.5	

Modeling Chemistry Unit 8 Mole Relationships Answers

Julia Schneider



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Table of Contents Modeling Chemistry Unit 8 Mole Relationships Answers

1. Understanding the eBook Modeling Chemistry Unit 8 Mole Relationships Answers
 - The Rise of Digital Reading Modeling Chemistry Unit 8 Mole Relationships Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 8 Mole Relationships Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 8 Mole Relationships Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 8 Mole Relationships Answers
 - Personalized Recommendations
 - Modeling Chemistry Unit 8 Mole Relationships Answers User Reviews and Ratings
 - Modeling Chemistry Unit 8 Mole Relationships Answers and Bestseller Lists

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 - Modeling Chemistry Unit 8 Mole Relationships Answers eBook Subscription Services
 - Modeling Chemistry Unit 8 Mole Relationships Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 8 Mole Relationships Answers Compatibility with Devices
 - Modeling Chemistry Unit 8 Mole Relationships Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 8 Mole Relationships Answers
 - Highlighting and Note-Taking Modeling Chemistry Unit 8 Mole Relationships Answers
 - Interactive Elements Modeling Chemistry Unit 8 Mole Relationships Answers
8. Staying Engaged with Modeling Chemistry Unit 8 Mole Relationships Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 8 Mole Relationships Answers
9. Balancing eBooks and Physical Books Modeling Chemistry Unit 8 Mole Relationships Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 8 Mole Relationships Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry Unit 8 Mole Relationships Answers
 - Setting Reading Goals Modeling Chemistry Unit 8 Mole Relationships Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry Unit 8 Mole Relationships Answers
 - Fact-Checking eBook Content of Modeling Chemistry Unit 8 Mole Relationships Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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