

Determination of K_a of a Weak Acid

(Show all data and calculations in your lab notebook)

Concentration of your standardized base 0.090 ± 0.0021

Table 1: Experimental Data

	Trial 1	Trial 2	Trial 3
Volume of unknown weak acid used, mL	34 mL	32 mL	33 mL
Volume of base required to reach equivalence point, mL	32.3 mL	30.2 mL	31.4 mL

Table 2: Experimental Results

	Trial 1	Trial 2	Trial 3	Mean \pm SD
Concentration of unknown weak acid, M	0.0846	0.0840	0.0847	0.084 \pm 0.00037
Volume of base required to reach half-equivalence point, mL	16.2	15.1	15.7	
pH at the half-equivalence point (from graph)	4.62	4.58	4.53	
pK_a of the unknown weak acid	4.62	4.58	4.53	
K_a of the unknown weak acid	2.3×10^{-5}	2.6×10^{-5}	2.9×10^{-5}	$2.6 \times 10^{-5} \pm 3 \times 10^{-6}$

Determination Of Ka Lab Report Answers

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