

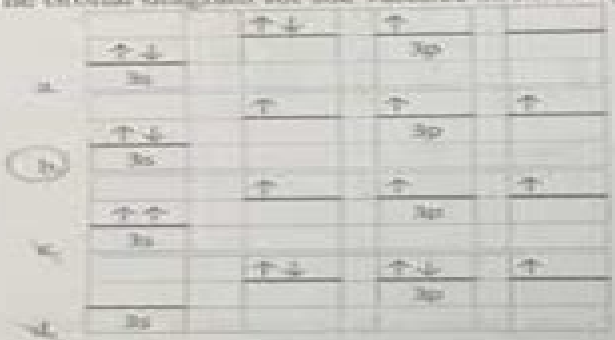
7.) Naturally-occurring copper exists as two isotopes: ^{63}Cu and ^{65}Cu , the masses of which are given below. Which of these isotopes, if either, is most abundant?

Isotope	Mass (amu)
^{63}Cu	62.929598
^{65}Cu	64.927730

- 35%
- a. More information is needed
 b. Neither, they are nearly-equal abundance
 c. ^{63}Cu is more abundant
 d. ^{65}Cu is more abundant

We can find the answer given the information and they would be abundant

8.) The orbital diagram for the valence shell of a phosphorous atom in its ground state is:



Section II – Compounds, Reactions, Mole Concept, Solutions, and Stoichiometry

9.) How many moles of xenon tetrafluoride correspond to 1.036 g of this compound?

- a. 2.5 mmol
 b. 1.5 mmol
 c. 500. mmol
 d. 5.0 mmol
 e. None of these

$1.036 \text{ g } \text{XeF}_4 \times \frac{1 \text{ mol}}{354.5 \text{ g}} = 0.0029 \text{ mol} \times 1000 \frac{\text{mmol}}{\text{mol}}$

10.) How many iodine atoms are present in 39.0 g of $\text{Ca}(\text{IO}_3)_2$?

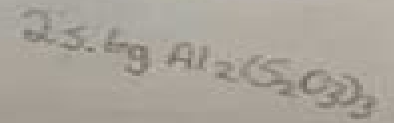
- a. 6.002×10^{23}
 b. 3.011×10^{23}
 c. 1.20×10^{24}
 d. 1.20×10^{22}
 e. None of these

$\text{Ca}(\text{IO}_3)_2 \quad 389.88 \text{ g mol}^{-1}$



$39.0 \text{ g} \times \frac{1 \text{ mol}}{389.88} \times \frac{2 \text{ mol I}}{1 \text{ mol}} = \text{mol I}$

$= 0.200$



$4.0 \times 10^{-1} \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}}$

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John Moore



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Table of Contents Honors Chemistry Semester Review Packet Answers

1. Understanding the eBook Honors Chemistry Semester Review Packet Answers
 - The Rise of Digital Reading Honors Chemistry Semester Review Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Honors Chemistry Semester Review Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Honors Chemistry Semester Review Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Honors Chemistry Semester Review Packet Answers
 - Personalized Recommendations
 - Honors Chemistry Semester Review Packet Answers User Reviews and Ratings

- Honors Chemistry Semester Review Packet Answers and Bestseller Lists
- 5. Accessing Honors Chemistry Semester Review Packet Answers Free and Paid eBooks
 - Honors Chemistry Semester Review Packet Answers Public Domain eBooks
 - Honors Chemistry Semester Review Packet Answers eBook Subscription Services
 - Honors Chemistry Semester Review Packet Answers Budget-Friendly Options
- 6. Navigating Honors Chemistry Semester Review Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Honors Chemistry Semester Review Packet Answers Compatibility with Devices
 - Honors Chemistry Semester Review Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Honors Chemistry Semester Review Packet Answers
 - Highlighting and Note-Taking Honors Chemistry Semester Review Packet Answers
 - Interactive Elements Honors Chemistry Semester Review Packet Answers
- 8. Staying Engaged with Honors Chemistry Semester Review Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Honors Chemistry Semester Review Packet Answers
- 9. Balancing eBooks and Physical Books Honors Chemistry Semester Review Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Honors Chemistry Semester Review Packet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Honors Chemistry Semester Review Packet Answers
 - Setting Reading Goals Honors Chemistry Semester Review Packet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Honors Chemistry Semester Review Packet Answers
 - Fact-Checking eBook Content of Honors Chemistry Semester Review Packet Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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