

REDOX REACTIONS.

NOTES...

Intro

~Many Chemical reactions involve transfer of electrons from one chemical substance to another. These electron-transfer reactions are termed as Oxidation-Reduction or Redox reactions.

Oxidation

(i) Loss of electron

- Example: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^{-}$

- Sub-points:

- Loss of hydrogen
- Gain of oxygen
- Gain of **electronegative** element OR loss of electropositive element

Reduction

i) Gain of electron

- Example: $\text{Cl}_2 + 2e^{-} \rightarrow 2\text{Cl}^{-}$

- Sub-points:

- Gain of hydrogen
- Loss of oxygen
- Loss of electronegative element OR gain of electropositive element

Redox reaction

The reaction in which oxidation as well as reduction takes place together is known as a redox reaction.

~REDOX REACTIONS IN TERMS OF ELECTRON TRANSFER (MODERN IDEA)

Oxidation process is the process in which one or more electrons are lost by an atom, ion, or molecule.

Example: $\text{Zn} \rightarrow \text{Zn}^{2+} + 2e^{-}$

Notes On Oxidation Reduction And Electrochemistry

VM Jensen



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Table of Contents Notes On Oxidation Reduction And Electrochemistry

1. Understanding the eBook Notes On Oxidation Reduction And Electrochemistry
 - The Rise of Digital Reading Notes On Oxidation Reduction And Electrochemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Notes On Oxidation Reduction And Electrochemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Notes On Oxidation Reduction And Electrochemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Notes On Oxidation Reduction And Electrochemistry
 - Personalized Recommendations
 - Notes On Oxidation Reduction And Electrochemistry User Reviews and Ratings

- Notes On Oxidation Reduction And Electrochemistry and Bestseller Lists
- 5. Accessing Notes On Oxidation Reduction And Electrochemistry Free and Paid eBooks
 - Notes On Oxidation Reduction And Electrochemistry Public Domain eBooks
 - Notes On Oxidation Reduction And Electrochemistry eBook Subscription Services
 - Notes On Oxidation Reduction And Electrochemistry Budget-Friendly Options
- 6. Navigating Notes On Oxidation Reduction And Electrochemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Notes On Oxidation Reduction And Electrochemistry Compatibility with Devices
 - Notes On Oxidation Reduction And Electrochemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Notes On Oxidation Reduction And Electrochemistry
 - Highlighting and Note-Taking Notes On Oxidation Reduction And Electrochemistry
 - Interactive Elements Notes On Oxidation Reduction And Electrochemistry
- 8. Staying Engaged with Notes On Oxidation Reduction And Electrochemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Notes On Oxidation Reduction And Electrochemistry
- 9. Balancing eBooks and Physical Books Notes On Oxidation Reduction And Electrochemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Notes On Oxidation Reduction And Electrochemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Notes On Oxidation Reduction And Electrochemistry
 - Setting Reading Goals Notes On Oxidation Reduction And Electrochemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Notes On Oxidation Reduction And Electrochemistry
 - Fact-Checking eBook Content of Notes On Oxidation Reduction And Electrochemistry
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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