



Step 6 & Final step

Chloride Content Calculation in Water

- Weight of sample (V_2) = 50 ml
- Titre value of water sample (V_1) = 19.3
- Titre value of blank (V_0) = 0.5

$$\begin{aligned} \text{Chloride, mg/l} &= \frac{(V_1 - V_0) \times N \times 35.45 \times 1000}{V_2} \\ &= \frac{(19.3 - 0.5) \times 0.01 \times 35.45 \times 1000}{50} \\ &= 127.94 \text{ mg/l} \end{aligned}$$

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Derek Midgley, Kenneth Torrance



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Analysis of the Chloride Ion Concentration in Aqueous Solutions by Potentiometric Titration Lyndsie Selwyn, Canadian Conservation Institute, 2001-01-01

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Potentiometry and Ion Selective Electrodes Alun Evans, 1987 For many chemists potentiometry is the first analytical procedure that they encounter during the use of an electrode to measure pH The advances in the field of ion selective electrodes now make it a valuable technique in the modern analytical laboratory

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Determination of Chloride in TATB. Progress Report, October--December 1975. [Potentiometric Analysis]., 1975 To determine total chloride content in TATB the sample was burned in a Parr Bomb and the chlorides formed were titrated potentiometrically using silver nitrate The electrode system is a chloride specific ion sensing element and a single junction reference Results on seven test samples agree generally within 10 percent with previous analyses by the Development Division Replicate analyses show precision of the method to be 10 relative percent Inorganic chloride is assayed by dissolving the TATB in concentrated sulfuric acid The HCl which is liberated is swept into dilute base by a nitrogen sweep and analyzed potentiometrically with the chloride specific electrode Results on three test samples agree within 5 percent with Development Division analyses with a 2 relative percent precision of analysis

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