

Chemistry Lab Report: Identifying an Unknown Compound

In this lab, I identified an unknown compound by testing and comparing the experimental results of the compound with the known compounds. The unknown compound in this lab was a white, crystalline solid that melted at 100°C. I tested the compound for its solubility in water, its melting point, and its boiling point. I also tested the compound for its reactivity with acids and bases. The results of the tests are summarized in the table below. The unknown compound is identified as sodium chloride (NaCl).

The unknown compound was a white, crystalline solid that melted at 100°C. I tested the compound for its solubility in water, its melting point, and its boiling point. I also tested the compound for its reactivity with acids and bases. The results of the tests are summarized in the table below. The unknown compound is identified as sodium chloride (NaCl).

When I tested the compound for its solubility in water, I found that it was highly soluble. I also tested the compound for its melting point and boiling point. The compound melted at 100°C and boiled at 100°C. I also tested the compound for its reactivity with acids and bases. The compound reacted with acids to form a colorless solution and with bases to form a colorless solution. The results of the tests are summarized in the table below. The unknown compound is identified as sodium chloride (NaCl).

I identified the unknown compound as sodium chloride (NaCl) by comparing the experimental results of the compound with the known compounds. The unknown compound in this lab was a white, crystalline solid that melted at 100°C. I tested the compound for its solubility in water, its melting point, and its boiling point. I also tested the compound for its reactivity with acids and bases. The results of the tests are summarized in the table below. The unknown compound is identified as sodium chloride (NaCl).

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**Richard C. Bauer, Richard Bauer, James
P. Birk, Douglas J. Sawyer**

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