

9/2/16

S. Sai Krishna

Date \_\_\_\_\_  
Page \_\_\_\_\_  
ClassmateThermochemistry

$$Q = \Delta E - W$$

$$Q = \Delta E - (-P\Delta V)$$

$$Q = \Delta E + P\Delta V$$

- At constant Volume.

$$Q = \Delta E$$

- At const. Pressure.

$$Q_p = \Delta E + P\Delta V$$

$$H = \Delta E + P\Delta V$$

$$H = \Delta E + \Delta nRT$$

$\Delta n =$  no. of moles of Prod  
- no. of moles of react.

→ For  $\Delta H$ .

- If  $\Delta H$  is -ve. reaction is exothermic.
- If  $\Delta H$  is +ve. reaction is endothermic.

Q. The heat of combustion of ethane at 17°C and const volume is  $-322.1 \text{ kJ}$ . Calculate the heat of combustion at constant pressure.



# Be Engineering Chemistry Notes Pdf 2016

**RJ Alexander**



**Be Engineering Chemistry Notes Pdf 2016:**

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Be Engineering Chemistry Notes Pdf 2016**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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