

## Solutions for Gravimetric Analysis Questions

### Check for Understanding 4.1

1. Determine the solubility of AgCl using  $K_{sp}$  for AgCl and a table of initial and equilibrium concentration terms.

	$\text{AgCl}(s)$	$\rightleftharpoons$	$\text{Ag}^+(aq)$	$+$	$\text{Cl}^-(aq)$	
init			0.0060		0	
equil			$0.0060 + x$		$x$	where $x$ = increase in $[\text{Cl}^-]$ = solubility of AgCl

at equil,  $K_{sp} = 1.8 \times 10^{-10} = [\text{Ag}^+][\text{Cl}^-] = (0.0060 + x)(x) = 0.0060x \rightarrow x = 3.0_0 \times 10^{-8} \text{ M}$

2.  $200. \text{ mL} \times \frac{3.0_0 \times 10^{-8} \text{ mol AgCl}}{1000 \text{ mL}} \times \frac{143.32 \text{ g AgCl}}{\text{mol AgCl}} = 8.6 \times 10^{-7} \text{ g AgCl}$

### Check for Understanding 4.2

1. Often the solubility of a slightly soluble salt is increased in acidic solution because the anion combines with  $\text{H}^+$  to form a weak acid. Removal of the anion shifts the equilibrium toward more dissolution of the salt. Since Cl would form a strong acid (HCl) that is ~100% ionized in solution, it is not removed from solution. Therefore, AgCl is not more soluble in acidic solution and the presence of  $\text{HNO}_3$  does not increase its solubility.

### Check for Understanding 4.3

1. First, determine the solubility of AgCl using  $K_{sp}$  for AgCl and a table of initial and equilibrium concentration terms.

	$\text{AgCl}(s)$	$\rightleftharpoons$	$\text{Ag}^+(aq)$	$+$	$\text{Cl}^-(aq)$	
init			0		0	
equil			$x$		$x$	

# Gravimetric Analysis Calculation Questions

**EW Minium**



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