

Molarity = Moles of solute / Liters of Solution (abbreviation = M)

Molality = Moles of solute / Kg of Solvent (abbreviation = m)

Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)

Mass Percent = mass of solute / mass of solution

1. How many moles of ethyl alcohol,  $C_2H_5OH$ , are present in 65 mL of a 1.5 M solution?

$$\frac{65 \text{ mL } C_2H_5OH}{1000 \text{ mL } C_2H_5OH} \times \frac{1 \text{ L } C_2H_5OH}{1 \text{ L } C_2H_5OH} \times \frac{1.5 \text{ mole } C_2H_5OH}{1 \text{ L } C_2H_5OH} = 0.098 \text{ moles } C_2H_5OH$$

2. How many liters of a 6.0 M solution of acetic acid  $CH_3COOH$ , contain 0.0030 moles of acetic acid?

$$\frac{0.0030 \text{ moles } CH_3COOH}{6.0 \text{ moles } CH_3COOH} \times \frac{1 \text{ L solution}}{1 \text{ L solution}} = 5.0 \times 10^{-4} \text{ L solution}$$

3. You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?

$$\frac{85 \text{ g KOH}}{56.11 \text{ g KOH}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{1 \text{ kg solvent}}{3.0 \text{ mole KOH}} \times \frac{1000 \text{ g solvent}}{1 \text{ kg solvent}} = 5.0 \times 10^2 \text{ g solvent}$$

$$5.0 \times 10^2 \text{ g solvent} + 85 \text{ g KOH} = 590 \text{ g solution}$$

4. If you dissolve 0.70 moles of HCl in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?

$$M = \frac{0.70 \text{ moles HCl}}{0.250 \text{ L solution}} = 2.8 \text{ M HCl}$$

5. A solution is prepared by adding 2.0 L of 6.0 M HCl to 500 mL of a 9.0 M HCl solution. What is the molarity of the new solution? (Remember, the volumes are additive)

$$\frac{2.0 \text{ L solution}}{1 \text{ L solution}} \times \frac{6.0 \text{ moles HCl}}{1 \text{ L solution}} = 12 \text{ moles HCl}$$

$$\frac{0.500 \text{ L solution}}{1 \text{ L solution}} \times \frac{9.0 \text{ moles HCl}}{1 \text{ L solution}} = 4.5 \text{ moles HCl}$$

$$12 \text{ moles HCl} + 4.5 \text{ moles HCl} = 17 \text{ moles}$$

$$2.0 \text{ L solution} + 0.500 \text{ L solution} = 2.5 \text{ L solution}$$

$$M = \frac{17 \text{ moles}}{2.5 \text{ L solution}} = 6.8 \text{ M HCl}$$

6. Convert the following Molarities to Normalities.

a. 2.5 M HCl = 2.5 N

b. 1.4 M  $H_2SO_4$  = 2.8 N

c. 1.0 M NaOH = 1.0 N

d. 0.5 M  $Ca(OH)_2$  = 1 N

# Molality Answers Key

**JR Anderson**



**Molality Answers Key:**

## **Molality Answers Key** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Molality Answers Key**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we will delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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