

## Chapter 9 Practice Problems: Chemical Names and Formulas

### Section 9.1: Naming Ions

- What is the charge on the ion typically formed by each element?
  - oxygen
  - iodine
  - sodium
  - aluminum
  - nickel, two electrons lost
  - magnesium
- How many electrons does the neutral atom gain or lose when each ion forms?
  - $\text{Cr}^{3+}$
  - $\text{P}^{3-}$
  - $\text{Li}^+$
  - $\text{Ca}^{2+}$
  - $\text{Cl}^-$
  - $\text{O}^{2-}$
- Name each ion. Identify each as a cation or an anion.
  - $\text{Sn}^{2+}$
  - $\text{Co}^{3+}$
  - $\text{Br}^-$
  - $\text{K}^+$
  - $\text{H}^-$
  - $\text{Mn}^{2+}$
- Write the formula (including the charge) for each ion. Use table 9.3 as if necessary.
  - carbonate ion
  - nitrite ion
  - sulfate ion
  - hydroxide ion
  - chromate ion
  - ammonium ion
- Name each of the following ions. Identify each as a cation or an anion.
  - $\text{CN}^-$
  - $\text{HCO}_3^-$
  - $\text{PO}_4^{3-}$
  - $\text{Cl}^-$
  - $\text{Ca}^{2+}$
  - $\text{SO}_3^{2-}$

### 9.2 Naming and Writing Formulas for Ionic Compounds:

- Write the formulas for these binary ionic compounds.
  - magnesium oxide
  - tin(II) fluoride
  - potassium iodide
  - aluminum chloride
  - sodium sulfide
  - ferric bromide
- Write the formulas for the compounds formed from these pairs of ions.
  - $\text{Ba}^{2+}$  and  $\text{Cl}^-$
  - $\text{Ag}^+$  and  $\text{I}^-$
  - $\text{Ca}^{2+}$  and  $\text{S}^{2-}$
  - $\text{K}^+$  and  $\text{Br}^-$
  - $\text{Al}^{3+}$  and  $\text{O}^{2-}$
  - $\text{Fe}^{2+}$  and  $\text{O}^{2-}$
- Name the following binary ionic compounds.
  - $\text{MnO}_2$
  - $\text{Li}_3\text{N}$
  - $\text{CaCl}_2$
  - $\text{SrBr}_2$
  - $\text{NiCl}_2$
  - $\text{K}_2\text{S}$
  - $\text{Fe}_3\text{N}_2$
  - $\text{SnCl}_4$
- Write formulas for the following ionic compounds.
  - sodium phosphate
  - magnesium sulfate
  - sodium hydroxide
  - potassium cyanide
  - ammonium chloride
  - potassium dichromate
- Write formulas for the following ionic compounds.
  - $\text{NH}_4^+$  and  $\text{SO}_4^{2-}$
  - $\text{K}^+$  and  $\text{NO}_3^-$
  - barium ion and hydroxide ion
  - lithium ion and carbonate ion

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