KINETICS Practice Problems and Solutions

Name: _____ AP Chemistry
Period: _____ Date: _____ Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
   a. This increases the $[H_2]$, which will increase the rate, but has no effect on k
   b. Due to Arrhenius equation, changing temperature changes the value of k.
   c. Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k.

6. Consider the reaction: $P_4 + 6\ H_2 \rightarrow 4\ PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/(L · s)
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

   a. What is the order with respect to: P_4 __0__.
      H_2 __1__.
   b. Write the rate law for this reaction. rate = $k[H_2]$
   c. Determine the value and units of the rate constant, k. plug and chug using the rate law & data from exp't 1 and solving for k, we get $k = 0.0427\ s^{-1}$

7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/(L · s)
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

   a. What is the order with respect to: SO_2 __2__.
      O_3 __0__.
   b. Write the rate law for this reaction. rate = $k[SO_2]^2[O_3]^0$
   c. Determine the value and units of the rate constant, k. plug and chug using the rate law & data from exp't 1 and solving for k, we get $k = 2.36\ mol^{-1}L^{-1} \cdot s^{-1}$

8. Consider the following mechanism.

$$A_2 + B_2 \rightarrow R + C \quad \text{(slow)}$$
$$A_2 + R \rightarrow C \quad \text{(fast)}$$

   a. Write the overall balanced chemical equation. $2\ A_2 + B_2 \rightarrow 2\ C$
   b. Identify any intermediates within the mechanism. R
   c. What is the order with respect to each reactant? A_2 1st; B_2 1st

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Table of Contents Chemical Kinetics Practice Problems And Answers

1. Understanding the eBook Chemical Kinetics Practice Problems And Answers
    - The Rise of Digital Reading Chemical Kinetics Practice Problems And Answers
    - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Kinetics Practice Problems And Answers
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Chemical Kinetics Practice Problems And Answers
    - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Kinetics Practice Problems And Answers
    - Personalized Recommendations
    - Chemical Kinetics Practice Problems And Answers User Reviews and Ratings
    - Chemical Kinetics Practice Problems And Answers and Bestseller Lists
5. Accessing Chemical Kinetics Practice Problems And Answers Free and Paid eBooks
    - Chemical Kinetics Practice Problems And Answers Public Domain eBooks
    - Chemical Kinetics Practice Problems And Answers eBook Subscription Services

- Chemical Kinetics Practice Problems And Answers Budget-Friendly Options
6. Navigating Chemical Kinetics Practice Problems And Answers eBook Formats
   - ePub, PDF, MOBI, and More
   - Chemical Kinetics Practice Problems And Answers Compatibility with Devices
   - Chemical Kinetics Practice Problems And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
   - Adjustable Fonts and Text Sizes of Chemical Kinetics Practice Problems And Answers
   - Highlighting and Note-Taking Chemical Kinetics Practice Problems And Answers
   - Interactive Elements Chemical Kinetics Practice Problems And Answers
8. Staying Engaged with Chemical Kinetics Practice Problems And Answers
   - Joining Online Reading Communities
   - Participating in Virtual Book Clubs
   - Following Authors and Publishers Chemical Kinetics Practice Problems And Answers
9. Balancing eBooks and Physical Books Chemical Kinetics Practice Problems And Answers
   - Benefits of a Digital Library
   - Creating a Diverse Reading Collection Chemical Kinetics Practice Problems And Answers
10. Overcoming Reading Challenges
   - Dealing with Digital Eye Strain
   - Minimizing Distractions
   - Managing Screen Time
11. Cultivating a Reading Routine Chemical Kinetics Practice Problems And Answers
   - Setting Reading Goals Chemical Kinetics Practice Problems And Answers
   - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Kinetics Practice Problems And Answers
   - Fact-Checking eBook Content of Chemical Kinetics Practice Problems And Answers
   - Distinguishing Credible Sources
13. Promoting Lifelong Learning
   - Utilizing eBooks for Skill Development
   - Exploring Educational eBooks
14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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