

## Study Card: Strong vs. Weak Acids and Bases

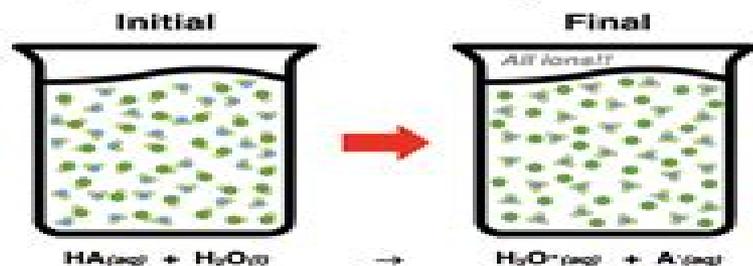
Acids and bases dissociate into ions in aqueous solutions. Whether an acid or a base is classified as strong or weak depends on the **degree of dissociation**.

**Strong Acids:** completely dissociate into ions when dissolved in water.



Goes to completion.

Equilibrium lies far to the right.  $K \gg \gg 1$ .

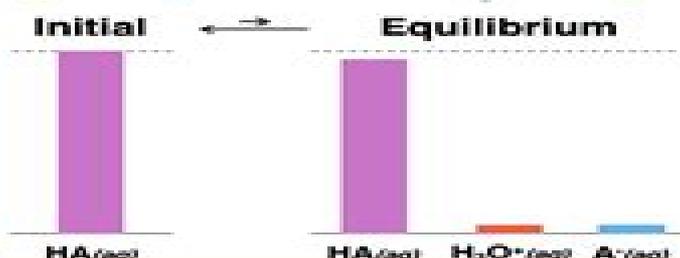
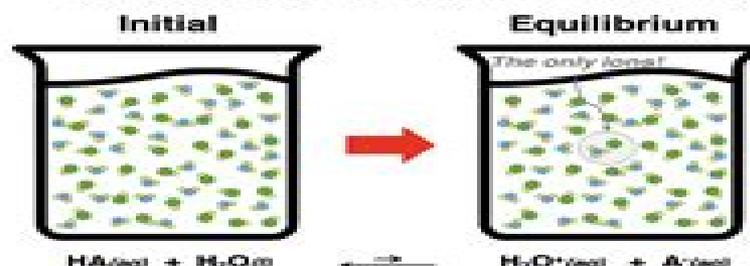


**Weak Acids:** partially dissociate into ions when dissolved in water.



Reversible; with a strong reverse reaction.

Equilibrium lies far to the left.  $K \ll \ll 1$ .



### Examples

Strong Acids	Strong Bases
Hydrochloric acid (HCl)	Lithium hydroxide (LiOH)
Sulfuric acid (H <sub>2</sub> SO <sub>4</sub> )	Sodium hydroxide (NaOH)
Nitric acid (HNO <sub>3</sub> )	Potassium hydroxide (KOH)
Hydrobromic acid (HBr)	Rubidium hydroxide (RbOH)
Hydroiodic acid (HI)	Cesium hydroxide (CsOH)
Perchloric acid (HClO <sub>4</sub> )	Calcium hydroxide (Ca(OH) <sub>2</sub> )
Chloric acid (HClO <sub>3</sub> )	Srtrontium hydroxide (Sr(OH) <sub>2</sub> )
	Barium hydroxide (Ba(OH) <sub>2</sub> )

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# 33 Strong Vs Weak Acids Answer

**RD Boyd**



### **33 Strong Vs Weak Acids Answer:**

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*A Course of Instruction in the General Principles of Chemistry* Arthur Amos Noyes, Miles Standish Sherrill, 1917

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*Journal of the Society of Chemical Industry* Society of Chemical Industry (Great Britain), 1923

**Elementary Chemistry** Harry Mann Gordin, 1913

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