



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - CH_4
 - H_2SO_4
- Write the formula for the following molecular compounds:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

- Write the molecular formula for each of the following compounds:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide
- Write the molecular formula for each of the following compounds:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. CH_4
- Show covalent bonding in carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Show covalent bonding in acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide



SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sublevel are represented by each of the following?
 - $2s$
 - $3p$
 - $4d$
 - $5f$
- Which orbital and sublevel are represented by each of the following?
 - $1s$
 - $2p$
 - $3d$
 - $4f$

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R Sandford



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