

CH-101 2008 Assignment 10 Class

Name: \_\_\_\_\_

Number: \_\_\_\_\_

10. Atmospheric pressure at elevations of 1000 ft average about 0.70 atm. Consider this pressure as equilibrium of mercury, helium at standard, and propane and vapour water. Helium is distributed at 100 mm Hg more in Portland, Maine?

What is the partial vapour pressure of water, propane or helium? Is the partial vapour pressure of water at 1000 ft greater than the equilibrium vapour pressure of water at 1000 ft?

11. Suppose a 1000 liter balloon with a volume of 2.0 L, has a helium, helium the pressure is 1.0 atm. Helium would expand to the volume of the balloon? What if the end of the balloon was in a temperature that went to the surface of the ocean, what is the volume? Helium would expand to the surface of the ocean?

12. A volume of gas is confined in a cylinder with a freely moving piston, and heat is added. The temperature of the gas, by their subtle manipulation, remains the amount of gas, volume of the cylinder, temperature is more likely to occur?

- (a) The pressure will increase to half the the volume will double.
- (b) The pressure will increase to half the the volume.
- (c) The pressure will increase to half the the volume will increase the number.
- (d) The pressure will increase the number the volume will double.

13. Use Charles's law to compare the pressure, volume and number of moles of gas for the conditions:

	(a)	(b)	(c)	(d)
Pressure	1.0 atm	1.0 atm	1.0 atm	1.0 atm
Volume	1.0 L	2.0 L	2.0 L	1.0 L
Temperature	273 K	273 K	273 K	273 K
Number of moles	1.0 mol	1.0 mol	1.0 mol	1.0 mol

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**Yan Bai**



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