

WORKSHEET
CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1
 Molecular formula of hydrocarbon or a mixture of hydrocarbon and oxygen

ANSWERS KEY

QUESTION 1. Calculate the empirical formula of hydrocarbon and mixture.

1. Determination of hydrocarbon hydrocarbon used by gas gas analysis method involves three reactions: $C_xH_y + O_2 \rightarrow CO_2 + H_2O$. From these, volume of CO_2 and H_2O are measured & are as follows:
 - (a) write the balanced chemical equation for the reaction: $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) compare the volume of CO_2 and H_2O as CO_2 and H_2O are 100 cm^3 and 150 cm^3 respectively
 - (c) deduce the volume of the hydrocarbon from the CO_2 and H_2O are 100 cm^3

2. State three experiments involved in volume analysis to determine the molecular formula of hydrocarbon. How hydrocarbon reacts with oxygen to form CO_2 and H_2O and how these are measured. These three steps of volume analysis are explained by the gas volume experiments:
 - (a) write the balanced chemical equation for the reaction: $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) determine the volume of CO_2 and H_2O as 100 cm^3 and 150 cm^3 respectively
 - (c) deduce the volume of the hydrocarbon from the CO_2 and H_2O are 100 cm^3

3. Hydrocarbon used in gas gas analysis is C_2H_6 . Write the chemical equation for the reaction: $C_2H_6 + O_2 \rightarrow CO_2 + H_2O$
 - (a) write the balanced chemical equation for the reaction: $C_2H_6 + O_2 \rightarrow CO_2 + H_2O$
 - (b) how many moles of C_2H_6 are produced by measurement of CO_2 and H_2O are 100 cm^3 and 150 cm^3 respectively
 - (c) how many moles of C_2H_6 are produced by measurement of CO_2 and H_2O are 100 cm^3 and 150 cm^3 respectively

QUESTION 2. Calculate the empirical formula of hydrocarbon.

1. From the data determine the empirical formula of hydrocarbon. From the volume analysis, determine the empirical formula of hydrocarbon. $C_xH_y + O_2 \rightarrow CO_2 + H_2O$. From the volume analysis, determine the empirical formula of hydrocarbon.
 - (a) write the balanced chemical equation for the reaction: $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) determine the empirical formula of hydrocarbon from the CO_2 and H_2O are 100 cm^3 and 150 cm^3 respectively

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