

1. A 33.7-g sample of HF is dissolved in water to give  $2.0 \times 10^2$  mL of solution. The concentration of the solution is:

- A) 1.68 M  
B) 0.34 M  
C) 0.17 M  
**D) 8.4 M**  
E) 16.9 M

$$[HF] = \frac{33.7g HF}{2.0 \times 10^2 mL} \times \frac{1000 mL}{1 L} \times \frac{1 mol HF}{20.01g HF} = 8.4 M$$

2. What mass of calcium chloride,  $CaCl_2$ , is needed to prepare 3.419 L of a 1.56 M solution?

- A) 243 g  
B) 5.33 g  
C) 50.6 g  
D) 111 g  
**E) 592 g**

$$\frac{3.419 L}{1} \times \frac{1.56 mol CaCl_2}{L} \times \frac{110.98g CaCl_2}{1 mol CaCl_2} = 592g CaCl_2$$

3. A 51.24-g sample of  $Ba(OH)_2$  is dissolved in enough water to make 1.205 liters of solution. How many mL of this solution must be diluted with water in order to make 1.000 liter of 0.1000 molar  $Ba(OH)_2$ ?

- A) 403 mL**  
B) 248 mL  
C) 24.8 mL  
D) 83.0 mL  
E) 331 mL

$$[Ba(OH)_2] = \frac{51.24g Ba(OH)_2}{1.205 L} \times \frac{1 mol Ba(OH)_2}{171.36g Ba(OH)_2} = 0.2481 M$$

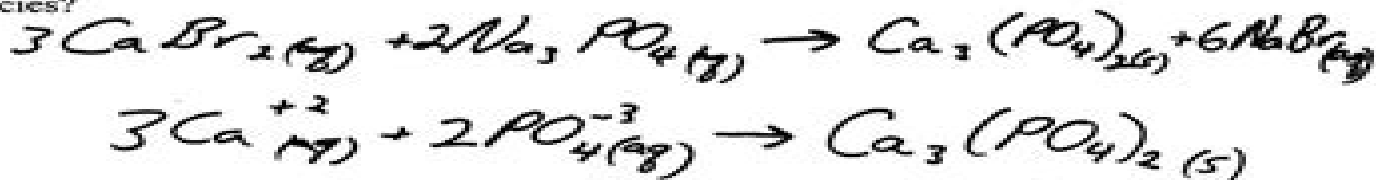
$$M_1 = 0.1000 M \quad M_2 = 0.2481 M$$

$$V_1 = 1.000 L \quad V_2 = ?$$

$$V_2 = \frac{0.1000 M \times 1.000 L}{0.2481 M} = 0.4031 L = 403 mL$$

4. The net ionic equation for the reaction of calcium bromide and sodium phosphate contains which of the following species?

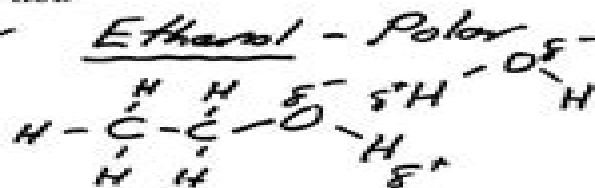
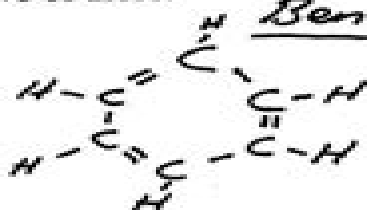
- A)  $Ca^{2+}(aq)$   
B)  $PO_4^{3-}(aq)$   
C)  $2Ca_3(PO_4)_2(s)$   
D)  $6NaBr(aq)$   
**E)  $3Ca^{2+}(aq)$**



5. Consider two organic molecules, ethanol and benzene. One dissolves in water and the other does not. Why?

- A) They have different molar masses.  
B) One is ionic, the other is not.  
C) One is an electrolyte, the other is not.  
**D) Ethanol contains a polar O—H bond, and benzene does not.**  
E) Two of these.

\* Know the general solubility rule of "Likes Dissolve Likes"



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**Thomas Evangelist**



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