

CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1** - Atomic Number
- H** - Atomic Symbol
- 1.008** - Atomic Weight



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For neutral atoms, $Z = +$ the number of electrons. The mass number, A , is the sum of the number of protons and the number of neutrons. Isotopes of an element have the same Z , different A . The atomic mass is the weighted average of isotopes. The atomic weight is the weighted average of isotopes for a given element.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number, n , is the number of shells. It is the number of nodes in the radial wave function. The number of orbitals in a shell is n^2 . The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$.

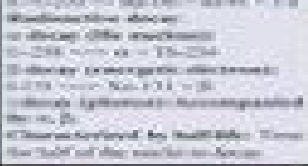
NUCLEAR CHEMISTRY

Nuclear reactions occur when nucleons (protons and neutrons) are rearranged. The energy released or absorbed is given by the mass defect. The mass defect is the difference between the mass of the reactants and the mass of the products. The energy released is given by $E = mc^2$.

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the many-electron atoms. The energy levels are split into subshells. The energy levels are split into subshells. The energy levels are split into subshells.

n	l	m _l	m _s	Orbitals
1	0	0	±1/2	1 orbital
2	0	0	±1/2	1 orbital
2	1	0, ±1	±1/2	3 orbitals
3	0	0	±1/2	1 orbital
3	1	0, ±1	±1/2	3 orbitals
3	2	0, ±1, ±2	±1/2	5 orbitals
4	0	0	±1/2	1 orbital
4	1	0, ±1	±1/2	3 orbitals
4	2	0, ±1, ±2	±1/2	5 orbitals
4	3	0, ±1, ±2, ±3	±1/2	7 orbitals



Atomic configuration: The arrangement of electrons in an atom. The Aufbau principle states that electrons fill orbitals from lowest to highest energy. The Pauli exclusion principle states that no two electrons can have the same set of quantum numbers. The Hund's rule states that electrons fill orbitals of equal energy one at a time, with parallel spins.

Block	Group	Element	Symbol	Atomic Number		
s-block	1	H	H	1		
		Li	Li	3		
	2	Na	Na	11		
		K	K	19		
	p-block	13	B	B	5	
			Al	Al	13	
		14	C	C	6	
			Si	Si	14	
		15	N	N	7	
			P	P	15	
		16	O	O	8	
			S	S	16	
		17	F	F	9	
			Cl	Cl	17	
		18	Ne	Ne	10	
			Ar	Ar	18	
		d-block	3	Sc	Sc	21
				Ti	Ti	22
4	V		V	23		
	Cr		Cr	24		
5	Mn		Mn	25		
	Fe		Fe	26		
6	Co		Co	27		
	Ni		Ni	28		
7	Cu		Cu	29		
	Zn		Zn	30		
f-block	7	La	La	57		
		Ce	Ce	58		
	8	Pr	Pr	59		
		Nd	Nd	60		
	9	Pm	Pm	61		
		Sm	Sm	62		
	10	Eu	Eu	63		
		Gd	Gd	64		
	11	Tb	Tb	65		
		Dy	Dy	66		
12	Ho	Ho	67			
	Er	Er	68			
13	Tm	Tm	69			
	Yb	Yb	70			
14	Lu	Lu	71			
	Hf	Hf	72			

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