

**Stoichiometry Worksheet 2**  
**Chemistry**

Name \_\_\_\_\_

Date \_\_\_\_\_ Hour \_\_\_\_\_

1. In the reaction :  $2\text{Al}_2\text{O}_{3(s)} \rightarrow 4\text{Al}_{(s)} + 3\text{O}_{2(g)}$ , the mole ratios are:

$\text{Al}_2\text{O}_3$  to Al \_\_\_\_\_  $\text{Al}_2\text{O}_3$  to  $\text{O}_2$  \_\_\_\_\_ Al to  $\text{O}_2$  \_\_\_\_\_  $\text{O}_2$  to Al \_\_\_\_\_

2. In the reaction:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ , the mole ratio of oxygen to water is: \_\_\_\_\_

3. In the reaction:  $\text{Ca} + \text{Cl}_2 \rightarrow \text{CaCl}_2$ , the mole ratio of chlorine to calcium is: \_\_\_\_\_

4. In the reaction:  $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$ , the mole ratio of zinc to sulfuric acid is:  
\_\_\_\_\_

5. The equation for the Haber process is represented by:



To convert 9.0 moles of hydrogen to ammonia would require how many moles of nitrogen?

6. In the equation:  $2\text{KClO}_{3(g)} \rightarrow 2\text{KCl}_{(s)} + 3\text{O}_{2(g)}$ , how many moles of oxygen are produced when 3.00 moles of  $\text{KClO}_3$  decompose completely?

7. For the reaction:  $\text{C}_{(s)} + 2\text{H}_{2(g)} \rightarrow \text{CH}_{4(g)}$ , how many moles of  $\text{H}_2$  are required to produce 10.0 moles of  $\text{CH}_4$ ?

# Solution Stoichiometry Worksheet Answers

**Michael Seilmaier**



**Solution Stoichiometry Worksheet Answers:**

## **Solution Stoichiometry Worksheet Answers** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the energy of words has be more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such is the essence of the book **Solution Stoichiometry Worksheet Answers**, a literary masterpiece that delves deep in to the significance of words and their effect on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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