

11-3 Practice Problems

limiting reactants - part 2

Name: _____

Date: _____

1. If 4.1 g of Cr is heated with 9.3 g of Cl_2 , what mass of CrCl_3 will be produced?



$$4.1\text{g} \times \frac{1\text{mole}}{52\text{g}} = 0.078\text{moles Cr}$$

$$9.3\text{g} \times \frac{1\text{mole}}{70\text{g}} = 0.133\text{moles Cl}_2$$

$$0.078\text{moles} \times \frac{3}{2} = 0.1182\text{moles Cl}_2 \text{ needed. Do I have it?}$$

2. What mass of SO_2 is produced from the reaction between 31.5 g of S_8 and 8.65 g of O_2 ?



3. What mass of SO_2 is produced from the reaction of 12.4 g of SO_2 and 3.45 g of O_2 ?



4. What mass of SO_2 is produced from the reaction of 58 g of SO_2 and 1.64 g of H_2O ?



11 3 Practice Answers

Michael Rajnik



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