

Science

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(Chapter - 4) (Structure of the Atom) (Exercises)

(Class - IX)

Question 1:

Compare the properties of electrons, protons and neutrons.

Answer 1:

Electron	Proton	Neutron
(i) Electron are present outside the nucleus of an atom.	(i) Proton are present in the nucleus of an atom.	(i) Neutron are present in the nucleus of an atom.
(ii) Electron are negatively charged	(ii) Proton are positively charged.	(ii) Neutron are neutral.
(iii) The mass of an electron is considered to negligible.	(iii) The mass of a proton is approximately 2000 times as the mass of an electron	(iii) The mass of neutron is nearly equal to the mass of a proton.

Question 2:

What are the limitations of J.J. Thomson's model of the atom?

Answer 2:

According to J.J. Thomson's model of an atom, an atom consists of a positively charged sphere with electrons embedded in it. However, it was later found that the positively charged particles reside at the center of the atom called the nucleus, and the electrons revolve around the nucleus.

Question 3:

What are the limitations of Rutherford's model of the atom?

Answer 3:

According to Rutherford's model of an atom, electrons revolve around the nucleus in fixed orbits. But, an electron revolving in circular orbits will not be stable because during revolution, it will experience acceleration. Due to acceleration, the electrons will lose energy in the form of radiation and fall into the nucleus. In such a case, the atom would be highly unstable and collapse.

Question 4:

Describe Bohr's model of the atom.

Answer 4:

Bohr's model of the atom Niels Bohr proposed the following postulates regarding the model of the atom.

- > Only certain orbits known as discrete orbits of electrons are allowed inside the atom.
- > While revolving in these discrete orbits, the electrons do not radiate energy. These discrete orbits or shells are shown in the following diagram.



The first orbit (i.e., for $n = 1$) is represented by letter K. Similarly, for $n = 2$, it is L - shell, for $n = 3$, it is M - shell and for $n = 4$, it is N - shell. These orbits or shells are also called energy levels.

Investment Science Chapter 4

Walter A. Rosenkrantz



Investment Science Chapter 4:

Introduction to Probability and Statistics for Science, Engineering, and Finance Walter A. Rosenkrantz, 2008-07-10

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Table of Contents Investment Science Chapter 4

1. Understanding the eBook Investment Science Chapter 4
 - The Rise of Digital Reading Investment Science Chapter 4
 - Advantages of eBooks Over Traditional Books
2. Identifying Investment Science Chapter 4
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Investment Science Chapter 4
 - User-Friendly Interface
4. Exploring eBook Recommendations from Investment Science Chapter 4
 - Personalized Recommendations
 - Investment Science Chapter 4 User Reviews and Ratings
 - Investment Science Chapter 4 and Bestseller Lists

5. Accessing Investment Science Chapter 4 Free and Paid eBooks
 - Investment Science Chapter 4 Public Domain eBooks
 - Investment Science Chapter 4 eBook Subscription Services
 - Investment Science Chapter 4 Budget-Friendly Options
6. Navigating Investment Science Chapter 4 eBook Formats
 - ePub, PDF, MOBI, and More
 - Investment Science Chapter 4 Compatibility with Devices
 - Investment Science Chapter 4 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Investment Science Chapter 4
 - Highlighting and Note-Taking Investment Science Chapter 4
 - Interactive Elements Investment Science Chapter 4
8. Staying Engaged with Investment Science Chapter 4
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Investment Science Chapter 4
9. Balancing eBooks and Physical Books Investment Science Chapter 4
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Investment Science Chapter 4
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Investment Science Chapter 4
 - Setting Reading Goals Investment Science Chapter 4
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Investment Science Chapter 4
 - Fact-Checking eBook Content of Investment Science Chapter 4
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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