

Name \_\_\_\_\_

Molecular Geometries Lab Report Sheet

Formula	Lewis Structure	No of electron domains/lig	Electron domain geometry	VSEPR Molecular geometry	Hybridization of central atom	Polar ?
CO <sub>2</sub>	<p>Valence 24</p> $\text{O}=\text{C}=\text{O}$ <p>2 charge clouds</p>	0	Linear 180°	Linear 	sp	NP
BF <sub>3</sub>	<p>Valence 32</p> $\text{F}=\text{B}-\text{F}$	0	Trigonal planar 120°		sp <sup>2</sup>	NP
SnCl <sub>2</sub>	<p>Valence 24</p> $\text{Cl}=\text{Sn}-\text{Cl}$	3	Bent 120°		sp <sup>2</sup>	P
BF <sub>4</sub> <sup>-</sup>	<p>Valence 40</p> $\left[ \text{F}-\text{B}-\text{F} \right]^-$	0	Tetrahedral 109.5°		sp <sup>3</sup>	P
AsH <sub>3</sub>	<p>Valence 24</p> $\text{H}-\text{As}-\text{H}$	4	Trigonal pyramidal 109.5°		sp <sup>3</sup>	P
H <sub>2</sub> S	<p>Valence 14</p> $\text{H}-\text{S}-\text{H}$	2	Bent 109.5°		sp <sup>3</sup>	NP
NH <sub>2</sub> <sup>-</sup>	<p>Valence 10</p> $\left[ \text{H}-\text{N}-\text{H} \right]^-$	2	Bent 109.5°		sp <sup>3</sup>	P
PCl <sub>4</sub>	<p>Valence 40</p> $\text{F}-\text{P}-\text{Cl}$	0	Trigonal bipyre 120 + 90		sp <sup>3</sup> d	P

# Molecular Shapes Lab Activity Answers

**JS Bruner**



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