

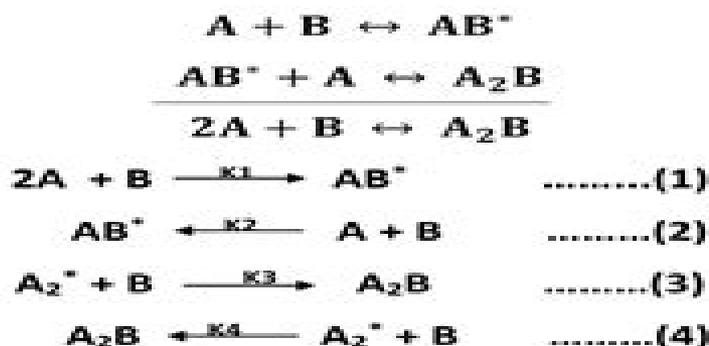
$$r_{A_2B} = \frac{K_3 \frac{K_1[A]^2}{2} [B]}{[K_3[B] + K_2]}$$

If K_2 is very small, $K_2 \cong 0$

$$r_{A_2B} = \frac{K_1 K_3 [A]^2 [B]}{K_3 [B]}$$

So, from above equation we can conclude that our first assumption is wrong so, mechanism also wrong.

Now, we can assume for second order reaction



From reaction (3) and (4) rate of formation of $r[A_2B]$,

$$r_{AB^*} = K_1[A][B] - K_2[AB^*] - K_3[AB^*][B] + K_4[A_2B] \quad \dots(5)$$

$$r[AB^*] = 0$$

$$0 = K_1[A][B] - K_2[AB^*] - K_3[AB^*][B] + K_4[A_2B]$$

$$[AB^*] = \frac{K_1[A][B] + K_4[A_2B]}{[K_3[A] + K_2]}$$

Put $[AB^*]$ value in equation (5),

$$r_{A_2B} = K_3 \left(\frac{K_1[A][B] + K_4[A_2B]}{[K_3[A] + K_2]} \right) [A] - K_4[A_2B]$$

Solution Manual Wangsness

Raffaela Di Napoli



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