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 molality = $\frac{\text{moles solute}}{\text{kilograms solvent}}$
 mathematical manipulation of molality is the
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 molality = $\frac{0.825 \text{ g} / 141.9579 \text{ g mol}^{-1}}{0.00581158 \text{ mol} / 0.00581158 \text{ mol}} = 0.4500 \text{ kg} / 0.0129 \text{ m}^3$
 mole fraction = $\frac{0.825 \text{ g} / 141.9579 \text{ g mol}^{-1}}{0.00581158 \text{ mol} / 0.00581158 \text{ mol} + 0.4500 \text{ kg} / 18.015 \text{ g mol}^{-1}}$
 = $0.0129 \text{ m}^3 / 0.00581158 \text{ mol} = 2.2024 \text{ hpo}$
 = $0.825 \text{ g} / 141.9579 \text{ g mol}^{-1} = 0.00581158 \text{ mol}$
 = $0.4500 \text{ kg} / 18.015 \text{ g mol}^{-1} = 24.97918401 \text{ mol}$

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$$\text{begin array l} \frac{\text{mass of a component}}{\text{mass of component in solution total mass of solution}} \times 100 \text{ end array}$$

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