

ANSWERS

Nuclear Chemistry Review Questions

1. Compare the mass, charge, and penetration power of alpha, beta and gamma radiation.

	Mass	Charge	Penetration
Alpha	4	+	Lowest
Beta	0	-	Middle
Gamma	0	0	Highest

2. In 5.49 seconds, 1.20g of argon-35 decays to leave only 0.15g. What is the half-life?

1.20g \rightarrow 0.60g \rightarrow 0.30g \rightarrow 0.15g, therefore it took 3 half-life periods.

$$5.49 \div 3 = 1.83 \text{ sec}$$

3. How many days does it take 16.0g of gold-198 to decay to 1.0g?

16.0g \rightarrow 8.0g \rightarrow 4.0g \rightarrow 2.0g \rightarrow 1.0g, therefore 4 half-life periods

From Ref Table N, 1 half-life = 2.69d

$$2.69\text{d} \times 4 = 10.76 \text{ days}$$

4. Why does an atom undergo radioactive decay?

Because it has an unstable nucleus

5. Compare and contrast fission and fusion. Why is fission used for electrical and fusion is not?

Fission=Division, the splitting of large unstable atoms(nucleus)

Fusion=joining together, small isotopes of H, creates way more energy

6. How is nuclear chemistry applied to medicine? What type of radioisotope must be used when it is injected into a person?

Radioisotopes are used to treat and diagnose illnesses

If injected they must have short half-lives and quickly removed from the body

7. Write out the nuclear decay of radium-226.



8. What is different about nuclear reactions when compared to a chemical reaction?

Nuclear reactions transform atoms into other atoms, chemical reactions involved the rearrangement of electrons

9. At what point do all isotopes of an element become radioactive?

Up to and including atomic # 83

Chapter 22 Review Nuclear Chemistry Answer Key

Darrell D. Ebbing



Chapter 22 Review Nuclear Chemistry Answer Key:

General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing, 1999 **DAT Prep Plus 2023-2024**
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Table of Contents Chapter 22 Review Nuclear Chemistry Answer Key

1. Understanding the eBook Chapter 22 Review Nuclear Chemistry Answer Key
 - The Rise of Digital Reading Chapter 22 Review Nuclear Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 22 Review Nuclear Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 22 Review Nuclear Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 22 Review Nuclear Chemistry Answer Key
 - Personalized Recommendations
 - Chapter 22 Review Nuclear Chemistry Answer Key User Reviews and Ratings
 - Chapter 22 Review Nuclear Chemistry Answer Key and Bestseller Lists
5. Accessing Chapter 22 Review Nuclear Chemistry Answer Key Free and Paid eBooks

- Chapter 22 Review Nuclear Chemistry Answer Key Public Domain eBooks
 - Chapter 22 Review Nuclear Chemistry Answer Key eBook Subscription Services
 - Chapter 22 Review Nuclear Chemistry Answer Key Budget-Friendly Options
6. Navigating Chapter 22 Review Nuclear Chemistry Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 22 Review Nuclear Chemistry Answer Key Compatibility with Devices
 - Chapter 22 Review Nuclear Chemistry Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 22 Review Nuclear Chemistry Answer Key
 - Highlighting and Note-Taking Chapter 22 Review Nuclear Chemistry Answer Key
 - Interactive Elements Chapter 22 Review Nuclear Chemistry Answer Key
 8. Staying Engaged with Chapter 22 Review Nuclear Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 22 Review Nuclear Chemistry Answer Key
 9. Balancing eBooks and Physical Books Chapter 22 Review Nuclear Chemistry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 22 Review Nuclear Chemistry Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 22 Review Nuclear Chemistry Answer Key
 - Setting Reading Goals Chapter 22 Review Nuclear Chemistry Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 22 Review Nuclear Chemistry Answer Key
 - Fact-Checking eBook Content of Chapter 22 Review Nuclear Chemistry Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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