

3 Scientific Measurement



QUANTIFYING MATTER



For students using the Foundation edition, assign problems 2-4, 7, 8, 10-16, 18-24.

3.1 Using and Expressing Measurements

Essential Understanding

In science, measurements must be accurate, precise, and written to the correct number of significant figures.



Reading Strategy

Venn Diagram A Venn diagram is a useful tool in visually organizing related information. A Venn diagram shows which characteristics the concepts share and which characteristics are unique to each concept.

As you read Lesson 3.1, use the Venn diagram to compare accuracy and precision.



EXTENSION Add the term *error* in the correct location in your Venn diagram. Then explain why you placed this term where you did.

Error is the difference between the measured value and the accepted value.

Lesson Summary

Scientific Notation Scientific notation is a kind of shorthand to write very large or very small numbers.

► Scientific notation always takes the form $(\text{a number} \geq 1 \text{ and } < 10) \times 10^n$.

Chapter 3 Scientific Measurement Practice Problems Answers

E Durkheim



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Foundations of College Chemistry, International Adaptation Morris Hein, Susan Arena, Cary Willard, 2023-05-16

Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students. It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence. The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need. This International Adaptation offers new and updated content with improved presentation of all course material. It builds on the strengths of previous editions including clear explanations and step by step problem solving. The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry. There is new and expanded coverage of polarizing power and polarizability, Fajans rules, collision number and mean free path, abnormal molecular masses and van t Hoff factor and applications of radioactivity. *Introduction to Chemical Principles* Howard Stephen Stoker, 1999. Chemistry for students who need full exposure to general chemistry but in compact one semester 17 chapter paperback format. Strong emphasis on problem solving with over 5000 problems in end of chapter material arranged in matched pairs. More real life applications added to this edition plus faces of chemistry.

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