

## Section 12.1

### OBJECTIVE 1

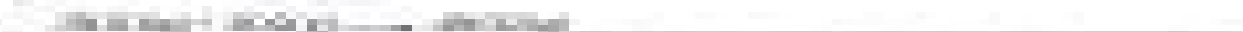
Write complete balanced redox reactions that are indicated.

1. \_\_\_\_\_
  - a. \_\_\_\_\_
    - (a)  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 + \text{H}^+ \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$
    - (b)  $\text{Mn}^{2+} + \text{H}_2\text{O}_2 + \text{OH}^- \rightarrow \text{MnO}_2 + \text{H}_2\text{O}$
    - (c)  $\text{Cr}^{3+} + \text{H}_2\text{O}_2 + \text{OH}^- \rightarrow \text{CrO}_4^{2-} + \text{H}_2\text{O}$
    - (d)  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 + \text{OH}^- \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$
  - b. \_\_\_\_\_
    - (a)  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 + \text{H}^+ \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$
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4. Write the net ionic equation for the following balanced equation:



a. Balance the equation:



b. Write the net ionic equation for  $\text{Mn}^{2+}$  and  $\text{H}_2\text{O}_2$ .

c. In a solution of  $\text{H}_2\text{O}_2$  there are many molecules of  $\text{H}_2\text{O}_2$  that have decomposed.

d. There are many molecules of  $\text{H}_2\text{O}_2$  that are not yet decomposed. How do you know?

e. The decomposition of  $\text{H}_2\text{O}_2$  is a reaction of  $\text{H}_2\text{O}_2$  and

# Chapter 9 Review Stoichiometry Section 1 Answers

**Ensheng Dong**



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